

**CHEMFILE MINI-GUIDE TO PROBLEM SOLVING****2. PLAN**

- *What step is needed to calculate the concentration of the diluted solution?*

Apply the principle that  $\text{volume}_1 \times \text{molarity}_1 = \text{volume}_2 \times \text{molarity}_2$ .

$$M_1 V_1 = M_2 V_2 \xrightarrow[\text{substitute each of the known quantities for its symbol and calculate}]{\begin{array}{l} \text{rearrange the equation} \\ M_1 V_1 = M_2 V_2 \\ \text{algebraically to solve for } M_2 \end{array}} M_2 = \frac{M_1 V_1}{V_2}$$

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

**3. COMPUTE**

Note: Even though molarity is moles per liter, you can use volumes in milliliters along with molarity whenever the units cancel.

$$M_2 = \frac{4.74 \text{ M} \times 50.00 \text{ mL}}{250.00 \text{ mL}} = 0.948 \text{ M}$$

**4. EVALUATE**

- *Are the units correct?*
- *Is the number of significant figures correct?*
- *Is the answer reasonable?*

Yes; molarity (mol/L) was required.  
 Yes; the number of significant figures is correct because the data had a minimum of three significant figures.  
 Yes; the computation is the same as  $4.74/5$ , which is a little less than 1.

**PRACTICE**

1. Complete the table below by calculating the missing value in each row.

	<b>Molarity of stock solution</b>	<b>Volume of stock solution</b>	<b>Molarity of dilute solution</b>	<b>Volume of dilute solution</b>	
a.	0.500 M KBr	20.00 mL	7 M KBr	100.00 mL	ans: 0.100 M
b.	1.00 M LiOH	? mL	0.075 M LiOH	500.00 mL	ans: 38 mL
c.	7 M HI	5.00 mL	0.0493 M HI	100.00 mL	ans: 0.986 M
d.	12.0 M HCl	0.250 L	1.8 M HCl	? L	ans: 1.7 L
e.	7.44 M NH <sub>3</sub>	? mL	0.093 M NH <sub>3</sub>	4.00 L	ans: 50. mL

# Chemfile Mini Guide To Problem Solving Answer Key

**Jicheng Xie**



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## **Table of Contents Chemfile Mini Guide To Problem Solving Answer Key**

1. Understanding the eBook Chemfile Mini Guide To Problem Solving Answer Key
  - The Rise of Digital Reading Chemfile Mini Guide To Problem Solving Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemfile Mini Guide To Problem Solving Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemfile Mini Guide To Problem Solving Answer Key
  - User-Friendly Interface

4. Exploring eBook Recommendations from Chemfile Mini Guide To Problem Solving Answer Key
  - Personalized Recommendations
  - Chemfile Mini Guide To Problem Solving Answer Key User Reviews and Ratings
  - Chemfile Mini Guide To Problem Solving Answer Key and Bestseller Lists
5. Accessing Chemfile Mini Guide To Problem Solving Answer Key Free and Paid eBooks
  - Chemfile Mini Guide To Problem Solving Answer Key Public Domain eBooks
  - Chemfile Mini Guide To Problem Solving Answer Key eBook Subscription Services
  - Chemfile Mini Guide To Problem Solving Answer Key Budget-Friendly Options
6. Navigating Chemfile Mini Guide To Problem Solving Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemfile Mini Guide To Problem Solving Answer Key Compatibility with Devices
  - Chemfile Mini Guide To Problem Solving Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemfile Mini Guide To Problem Solving Answer Key
  - Highlighting and Note-Taking Chemfile Mini Guide To Problem Solving Answer Key
  - Interactive Elements Chemfile Mini Guide To Problem Solving Answer Key
8. Staying Engaged with Chemfile Mini Guide To Problem Solving Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemfile Mini Guide To Problem Solving Answer Key
9. Balancing eBooks and Physical Books Chemfile Mini Guide To Problem Solving Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemfile Mini Guide To Problem Solving Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemfile Mini Guide To Problem Solving Answer Key
  - Setting Reading Goals Chemfile Mini Guide To Problem Solving Answer Key
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemfile Mini Guide To Problem Solving Answer Key
  - Fact-Checking eBook Content of Chemfile Mini Guide To Problem Solving Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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