

Chapter 17: Chemical Equilibrium

Worksheet 17.1: K_c & K_p

1. Consider the following reaction:

Reaction

- 1. Write the expression for the equilibrium constant, K_c , for the reaction.
- 2. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.
- 3. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.
- 4. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.
- 5. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.
- 6. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.
- 7. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.
- 8. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.
- 9. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.
- 10. Calculate the value of K_c at 25°C if the concentrations are: $[N_2] = 0.10 \text{ mol L}^{-1}$, $[O_2] = 0.10 \text{ mol L}^{-1}$, and $[NO] = 0.040 \text{ mol L}^{-1}$.

Answers

- 1. $K_c = \frac{[NO]^2}{[N_2][O_2]}$
- 2. $K_c = \frac{(0.040)^2}{(0.10)(0.10)} = 0.16$
- 3. $K_c = \frac{(0.040)^2}{(0.10)(0.10)} = 0.16$
- 4. $K_c = \frac{(0.040)^2}{(0.10)(0.10)} = 0.16$
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- 9. $K_c = \frac{(0.040)^2}{(0.10)(0.10)} = 0.16$
- 10. $K_c = \frac{(0.040)^2}{(0.10)(0.10)} = 0.16$

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