

Week 10: Problem

1. 2018, 2019, 2020

Answer: Diagram

Introduction: 10 min

This standard integrates the first and second laws of thermodynamics developed in

classical thermodynamics.

It emphasizes: Equilibrium State, State Functions and processes between states of a closed system.

What is said

States are represented as points in a thermodynamic state space and process the evolution in this space is traced.

1. In any process, the quantity "heat" Q is not defined. Heat is associated with the flow of energy.

which includes a system both as a whole

processes, processes that range, and thermodynamic quantities.

2. The quantities heat, work and enthalpy are not state functions. The actual path of a process is significant in their values.

Thermodynamic Quantities:

Internal energy

Enthalpy

Heat capacity

Temperature

3. Does the "heat" Q include what flows into, out of, or through a thermodynamic state as they

the state evolves? Heat is the energy that is transferred between a system and its surroundings.

4. Entropy: Is it a state function? If not, what is the state function? Is it a state function? Is it a state function?

Entropy is a state function. It is defined by the integral of the heat Q divided by the temperature T .

5. Entropy: What is the state function that is a measure of the degree of order? Entropy is a state function that is a measure of the degree of order.

6. Entropy: Entropy is a state function, only in the thermodynamic

equilibrium state. Entropy is a state function only in the thermodynamic

equilibrium state. Entropy is a state function only in the thermodynamic

equilibrium state.

7. Entropy is a state function

8. Entropy is a state function

9. Entropy is a state function

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