

Chemistry - Chapter 9 Assessment Answers (Pages 281-282)

42. a) 2- b) 1+ c) 1- d) 3+

43. a) 2+ b) 2+ c) 3+ d) 1+

44. a) barium ion b) iodide ion c) silver ion d) mercury(II) ion

45. Cyanide,  $\text{CN}^-$  and hydroxide,  $\text{OH}^-$

46. a) hydroxide ion b) lead(IV) ion c) sulfate ion d) oxide ion

47. zero; A compound is electrically neutral.

48. The symbols for the cation and anion are written and the charges are balanced with subscripts. The name of the cation is followed by the name of the anion.

49. Determine the charge of the anion then work the formula backwards to find the charge of the transition metal cation needed to give a net charge of zero for the formula unit.

50. The symbols for the cation and anion are written and the charges are balanced with subscripts. Parentheses are used around the polyatomic ion if a subscript is needed. The name of the cation is followed by the name of the anion.

51. a and b

52. When more than a single polyatomic ion is needed to balance the formula.

53.  $\text{NH}_4\text{NO}_3$ , ammonium nitrate;  $(\text{NH}_4)_2\text{CO}_3$ , ammonium carbonate;  $\text{NH}_4\text{CN}$ , ammonium cyanide;  $(\text{NH}_4)_3\text{PO}_4$ , ammonium phosphate;  $\text{Sn}(\text{NO}_3)_4$ , tin(IV) nitrate;  $\text{Sn}(\text{CO}_3)_2$ , tin(IV) carbonate;  $\text{Sn}(\text{CN})_4$ , tin(IV) cyanide;  $\text{Sn}_2(\text{PO}_4)_4$ , tin(IV) phosphate;  $\text{Fe}(\text{NO}_3)_3$ , iron(III) nitrate;  $\text{Fe}_2(\text{CO}_3)_3$ , iron(III) carbonate;  $\text{Fe}(\text{CN})_3$ , iron(III) cyanide;  $\text{FePO}_4$ , iron(III) phosphate;  $\text{Mg}(\text{NO}_3)_2$ , magnesium nitrate;  $\text{MgCO}_3$ , magnesium carbonate;  $\text{Mg}(\text{CN})_2$ , magnesium cyanide;  $\text{Mg}_3(\text{PO}_4)_2$ , magnesium phosphate

54. two nonmetals

55. a) tri- b) mono- c) di- d) hexa- e) penta- f) tetra-

56. For formulas, write the correct symbols for each element with a subscript corresponding to the prefix before each element in the name. For naming, name each element in the order given. Use the subscript to determine the prefixes before each element in the name. The name ends in *-ide*.

57. a)  $\text{BCl}_3$  b) dinitrogen pentoxide c)  $\text{N}_2\text{H}_4$  d) carbon tetrachloride

58. a) hydrochloric acid b)  $\text{H}_2\text{SO}_4$  c) nitric acid d)  $\text{HC}_2\text{H}_3\text{O}_2$

# Mastering Chemistry Chapter 9 Answers

**M Walker**



## Mastering Chemistry Chapter 9 Answers:

*Chemistry* John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science **Mastering Chemistry** Jason Chin, Vault Reports (Firm), Job Vault, 1998 The periodic table never looked so good From the properties of matter and mole calculations to the kinetics of reactions and enthalpy change this guide demystifies the complexities of chemistry through easy to follow charts graphs and sample problems EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS CHANG, 2013-01-07 EBOOK GENERAL CHEMISTRY THE ESSENTIAL CONCEPTS Mastering Chemistry Gerald W. Gibson, 1975 Stewart's Textbook of Acid-Base John A Kellum, Paul WG Elbers, 2009 Rev ed of How to understand acid base c1981 General Chemistry for Engineers Jeffrey Gaffney, Nancy Marley, 2017-11-13 General Chemistry for Engineers explores the key areas of chemistry needed for engineers This book develops material from the basics to more advanced areas in a systematic fashion As the material is presented case studies relevant to engineering are included that demonstrate the strong link between chemistry and the various areas of engineering Serves as a unique chemistry reference source for professional engineers Provides the chemistry principles required by various engineering disciplines Begins with an atoms first approach building from the simple to the more complex chemical concepts Includes engineering case studies connecting chemical principles to solving actual engineering problems Links chemistry to contemporary issues related to the interface between chemistry and engineering practices Chemistry John S. Phillips, Cheryl Wistrom, 2000 A Problem-Solving Approach to Aquatic Chemistry James N. Jensen, 2022-12-20 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter s main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and

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Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Daluis Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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more detailed explanations in solutions and chapter outlines *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 1982-01-01 *Study Guide to Accompany Basics for Chemistry* is an 18 chapter text designed to be used with *Basics for Chemistry* textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Elements of Inorganic Chemistry**

James Henry Shepard, 1885 **Mastering Organic Chemistry** Gerald W. Gibson, 1979 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Mastering Pharmacy Licensure and Certification David C. Kosegarten, Douglas J. Pisano, F. Randy Vogenberg, 2000 Integrating knowledge from the core pharmacy basic disciplines this guide presents over 100 intriguing case studies Each study is followed by ten to 15 questions designed to prepare students for the pharmacy board exam Five of the questions are clinically oriented and focus on the 200 commonly used drugs

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