

1. Formaldehyde (CH₂O) is a colorless, pungent gas that is widely used in the chemical industry and as a preservative in formalin solutions.

Physical Properties

At room temperature, formaldehyde is a colorless gas with a strong, irritating odor.

Chemical Properties

Formaldehyde is highly reactive and can undergo various chemical reactions.

Reactions with Water

Formaldehyde reacts with water to form formalin, a solution of formaldehyde in water.

Reaction with Alcohols

Formaldehyde reacts with alcohols to form hemiacetals.

Reaction with Aldehydes and Ketones

Formaldehyde reacts with other aldehydes and ketones to form acetals.

Reaction with Amines

Formaldehyde reacts with amines to form imines.

Formaldehyde reacts with primary amines to form imines.

Reaction with Phenols

Formaldehyde reacts with phenols to form phenol-formaldehyde resins.

Formaldehyde reacts with phenols to form phenol-formaldehyde resins.

Formaldehyde reacts with phenols to form phenol-formaldehyde resins.

Reaction with Carbohydrates

Formaldehyde reacts with carbohydrates to form cross-linked polymers.

Formaldehyde reacts with carbohydrates to form cross-linked polymers.

Reaction with Proteins

Formaldehyde reacts with proteins to form cross-linked polymers.

Formaldehyde reacts with proteins to form cross-linked polymers.

Reaction with Nucleic Acids

Formaldehyde reacts with nucleic acids to form cross-linked polymers.

Formaldehyde reacts with nucleic acids to form cross-linked polymers.

Reaction with Cellulose

Formaldehyde reacts with cellulose to form cross-linked polymers.

Bouncers Script Pdf

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