

Class 10th Chemistry Practical Copy

Lab No. _____ Date _____

of solution can be observed only by comparing the hydrogen ion concentration of solutions. Then the concentration of weak solution and an equal amount of strong solution has been compared with weak solution. Spent of pH value. The pH value was found to be _____

Weak solution: $[H^+] = 10^{-4} M$, pH = 4.00
 Strong solution: $[H^+] = 10^{-2} M$, pH = 2.00
 Neutral solution: $[H^+] = 10^{-7} M$, pH = 7.00

1. Solution having a pH of 2 will be a weak acid solution with pH value above 7 are basic in nature and solution with pH value below 7 are acids.

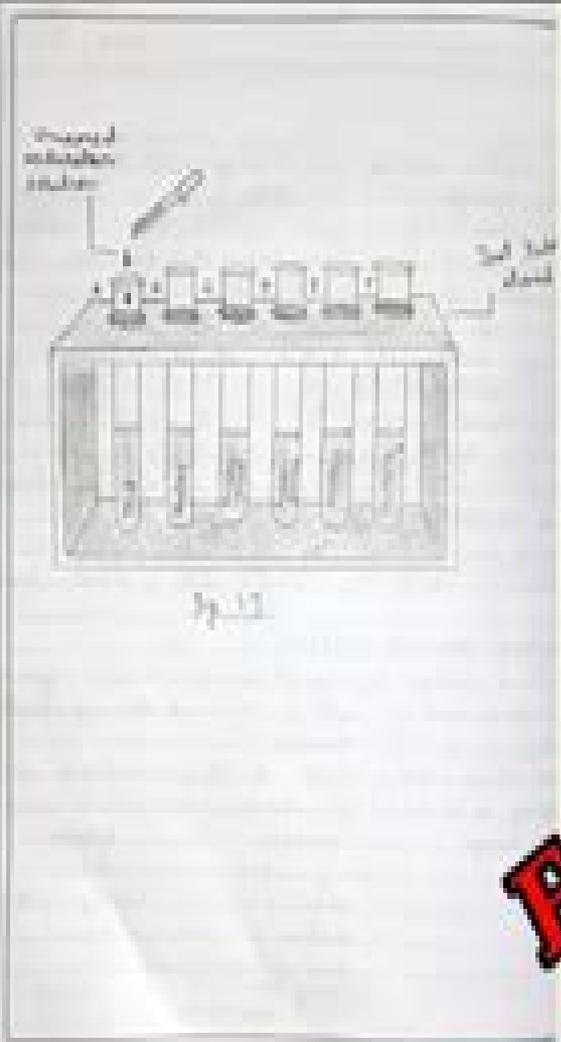
2. pH of a solution can be determined using a pH meter or using an indicator with universal indicator.

3. Universal Indicator Solution

It is a solution of several indicators having different pH ranges and is used to determine the approximate value of pH of a solution.

4. The colour that indicates the pH of a solution after adding a universal indicator is:

pH range	Indicator	Colour
1-2	Strong acid	Red
3-4	Red	Orange/Yellow
5-6	Yellow	Green
7-8	Blue	Blue
9-10	Strong base	Dark Blue



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