

16. ^{137}I has a half-life of 8.04 days. Assuming you start with a 1.25 mg sample of ^{137}I , how much will remain after 13.8 days?

- A) 0.433 mg
 B) 0.268 mg
 C) 0.422 mg
 D) 0.388 mg

$$k = \frac{\ln 2}{t_{1/2}} = \frac{\ln 2}{8.04 \text{ days}} = 0.0862 \text{ days}^{-1}$$

$$\ln(N_t) = -kt + \ln(N_0)$$

$$\ln(N_t) = -(0.0862 \text{ days}^{-1})(13.8 \text{ days}) + \ln(1.25 \text{ mg}) = -0.823$$

$$N_t = e^{-0.823} = \mathbf{0.433 \text{ mg}}$$

17. Which one of the following devices converts radioactive radiation to light for detection?

- A) Geiger counter
 B) photomultiplier tube
 C) scintillation counter
 D) none of these

18. Which one of the following is true?

- A) Most spontaneous nuclear reactions are exothermic.
 B) Some spontaneous nuclear reactions are endothermic.
 C) All spontaneous nuclear reactions are exothermic.
 D) There is no relationship between exothermicity and spontaneity in nuclear reactions.

19. Binding energy per nucleon is greatest for nuclei with...

- A) large mass numbers
 B) intermediate mass numbers
 C) small mass numbers

20. When two atoms of ^{235}U are fused to form one atom of ^{238}U , the total energy involved is $3.83 \times 10^{-11} \text{ J}$. What is the total change in mass for this reaction? ($c = 3.00 \times 10^8 \text{ m/s}$)

- A) $4.26 \times 10^{-17} \text{ g}$
 B) $4.26 \times 10^{-16} \text{ g}$
 C) $4.26 \times 10^{-18} \text{ g}$
 D) $4.26 \times 10^{-19} \text{ g}$

$$E = mc^2 \quad \Rightarrow \quad m = \frac{E}{c^2} = \frac{3.83 \times 10^{-11} \text{ J}}{(3.00 \times 10^8 \text{ m/s})^2} = 4.26 \times 10^{-16} \text{ kg} = \left(\frac{1000 \text{ g}}{1 \text{ kg}}\right) = \mathbf{4.26 \times 10^{-16} \text{ g}}$$

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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2002-01-04 This third edition continues to innovate by providing students with an integrated and modern approach to the subject The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors Also of benefit is the just in time presentation of key content only providing details once they are needed While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons The whole is backed by a range of supplements including a new illustration program a tutorial CD interactive learningware an extensive Web CT component an instructor's resource CD and a solution CD

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Engineering News ,1963 **El-Hi Textbooks in Print** ,1984 *ERDA Energy Research Abstracts* United States. Energy Research and Development Administration,1976-10 *Hands-On Chemistry Activities with Real-Life Applications* Norman Herr,James Cunningham,1999-01-13 This comprehensive collection of over 300 intriguing investigations including demonstrations labs and other activities uses everyday examples to make chemistry concepts easy to understand It is part of the two volume PHYSICAL SCIENCE CURRICULUM LIBRARY which consists of Hands On Physics Activities With Real Life Applications and Hands On Chemistry Activities With Real Life Applications

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Table of Contents Study Nuclear Chemistry Answer Key

1. Understanding the eBook Study Nuclear Chemistry Answer Key
 - The Rise of Digital Reading Study Nuclear Chemistry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Nuclear Chemistry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Nuclear Chemistry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Nuclear Chemistry Answer Key
 - Personalized Recommendations
 - Study Nuclear Chemistry Answer Key User Reviews and Ratings
 - Study Nuclear Chemistry Answer Key and Bestseller Lists
5. Accessing Study Nuclear Chemistry Answer Key Free and Paid eBooks

- Study Nuclear Chemistry Answer Key Public Domain eBooks
 - Study Nuclear Chemistry Answer Key eBook Subscription Services
 - Study Nuclear Chemistry Answer Key Budget-Friendly Options
6. Navigating Study Nuclear Chemistry Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Study Nuclear Chemistry Answer Key Compatibility with Devices
 - Study Nuclear Chemistry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Study Nuclear Chemistry Answer Key
 - Highlighting and Note-Taking Study Nuclear Chemistry Answer Key
 - Interactive Elements Study Nuclear Chemistry Answer Key
8. Staying Engaged with Study Nuclear Chemistry Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Nuclear Chemistry Answer Key
9. Balancing eBooks and Physical Books Study Nuclear Chemistry Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Nuclear Chemistry Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Nuclear Chemistry Answer Key
- Setting Reading Goals Study Nuclear Chemistry Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Nuclear Chemistry Answer Key
- Fact-Checking eBook Content of Study Nuclear Chemistry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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