

**WORKSHEET**  
**CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1**  
 Molecular formula of hydrocarbons and the structural formula of hydrocarbons

**ANSWERS KEY**

**QUESTION 1. Calculate the molecular formula of hydrocarbons and draw their structure.**

1. Molecular formula of hydrocarbon is calculated as given by the given carbon and hydrogen atoms respectively.  $C_{10}H_{22}$   $\Rightarrow$   $C_{10}H_{22}$  or  $C_{10}H_{2n+2}$  (This means presence of 10 carbon and hydrogen atoms)
  - (a) write the molecular structural equation for this compound:  
 $C_{10}H_{22} \Rightarrow$  or  $C_{10}H_{2n+2}$
  - (b) calculate the number of hydrogen atoms in molecule as  $2n+2$  where  $n = 10$  or
  - (c) draw the structure of the compound from the given molecular formula.
  
2. Write the structural formulae of given hydrocarbons to determine the carbon-carbon (C-C) bonding arrangement. The hydrocarbons are  $C_4H_{10}$  and  $C_5H_{12}$ . Draw the structural formulae of hydrocarbons and name them. These show presence of carbon-carbon single bond as indicated by 1 line (C-C single bond).
  - (a) write the molecular structural equation for this compound:  
 $C_4H_{10}$  or  $C_4H_{2n+2}$   $\Rightarrow$   $C_4H_{10}$  or  $C_4H_{2n+2}$
  - (b) calculate the number of hydrogen atoms in molecule as  $2n+2$  where  $n = 4$  or
  - (c) draw the structure of carbon-carbon chain as indicated in  $C_4H_{10}$  or  $C_4H_{2n+2}$ .
  
3. Hydrocarbons are given in two cases (carbon and hydrogen) respectively as the compound is  $C_3H_8$  or  $C_3H_{2n+2}$ 
  - (a) write the molecular structural equation for this compound:  
 $C_3H_8$  or  $C_3H_{2n+2}$   $\Rightarrow$   $C_3H_8$  or  $C_3H_{2n+2}$
  - (b) draw the structure of this compound as indicated by the molecular formula  $C_3H_8$  or  $C_3H_{2n+2}$ .
  - (c) draw the structure of hydrocarbon as indicated by the given molecular formula  $C_3H_8$  or  $C_3H_{2n+2}$ .

**QUESTION 2. Calculate the molecular formula of hydrocarbons and draw their structure.**

1. The total number of carbon-carbon bonds for calculating the C-C bond is as given (carbon-carbon bonding arrangement) as given carbon number (C<sub>n</sub>) and the given hydrogen atoms  $C_4H_{10}$  or  $C_4H_{2n+2}$   $\Rightarrow$   $C_4H_{10}$  or  $C_4H_{2n+2}$  (This means presence of 4 carbon and hydrogen atoms)
  - (a) write the molecular structural equation for this compound:  
 $C_4H_{10}$  or  $C_4H_{2n+2}$   $\Rightarrow$   $C_4H_{10}$  or  $C_4H_{2n+2}$
  - (b) calculate the number of hydrogen atoms in the compound as  $2n+2$ .

# Chem 21 Lab Answers

**Thirumalaisamy P.  
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