

Thermodynamic processes:

Process in which state of thermodynamic variables like P, V, T changes are called thermodynamic processes.

Given fixed number of moles / closed system ($n = \text{constant}$) under ideal conditions is assumed ($PV = nRT$ is applicable)

Constant volume process / Isochoric process / Isometric process:

Gas is confined in a piston cylinder arrangement having fixed piston.

$$V = \text{constant}$$

$$PV = nRT \Rightarrow \frac{P}{T} = \left(\frac{nR}{V}\right) = \text{const.}$$

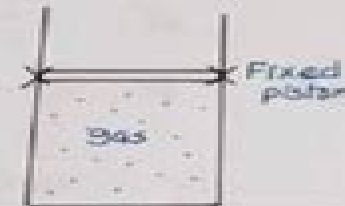
Process equation: $\frac{P}{T} = \text{constant}$

Between two states (P_1, V, T_1) and (P_2, V, T_2)

$$\text{one can use: } \frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \left\{ T_1, T_2 \text{ must be in kelvin} \right\}$$

Example: A given mass of an ideal gas is confined in a vessel of fixed volume. At 27°C pressure of the gas is found to be 2.4 atm . Find its pressure if the gas is heated to 52°C .

$$\begin{aligned} \text{Sol: } P_1 &= 2.4 \text{ atm} & , & T_1 = 273 + 27 = 300 \text{ K} \\ P_2 &= ? & , & T_2 = 273 + 52 = 325 \text{ K} \\ \frac{P_1}{T_1} &= \frac{P_2}{T_2} \Rightarrow P_2 = \frac{T_2}{T_1} \times P_1 = \frac{325}{300} \times 2.4 = 2.6 \text{ atm} \\ & & & \text{(answer)} \end{aligned}$$



Chemical Engineering Lecture Notes

M Walker



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Chemical Engineering Education, 2002 **Catalog** South Dakota School of Mines and Technology, 1927
Statistical Modelling of Molecular Descriptors in QSAR/QSPR Matthias Dehmer, Kurt Varmuza, Danail

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