

# CHAPTER # 1

## STOICHIOMETRY

**Q1. Define stoichiometry.**

**Ans: Stoichiometry:**

The study of relative amounts of substances involved in a chemical reaction is called Stoichiometry. Such phenomenon is studied through the knowledge of Stoichiometry (Greek word *Stolcheion* means element and *metry* means measurement).

**Importance of Stoichiometry:**

Stoichiometry is essential when quantitative information about a chemical reaction is required. Moreover it is important to predict yields of chemical products.

**Q2. Explain the significance of stoichiometry with the help of example.**

OR

**How will you explain law of conservation of mass in the case of combustion of hydrogen fuel in rockets?**

**Ans: Combustion of hydrogen fuel in rockets:**

1. Consider a rocket manufacturer uses liquid hydrogen as a fuel. He may have to determine how much fuel is necessary for a particular flight. Hydrogen burns in oxygen (of air) to produce water.



It states that

- (i) Two moles of Hydrogen react with one mole of oxygen to form two moles of steam.
  - (ii) Two molecules of Hydrogen react with one molecule of oxygen to produce two molecules of steam.
  - (iii) Four grams of hydrogen react with thirty-two grams of oxygen to produce thirty-six grams of water. Here the total mass of reactants is equal to the total mass of products. Thus it confirms the Law of conservation of mass.
2. Another example is the reaction taking place in a gas barbecue. This is the example of combustion to form carbon dioxide and water. The balanced chemical reaction is



**Q3. What do you understand by the term Mole?**

**Ans: Mole:**

The atomic mass, formula mass and molecular mass of a substance expressed in grams is called Mole.

$$\text{Number of moles} = \frac{\text{mass in grams}}{\text{molecular mass}}$$

**Example:** One mole of O = 16 g

# Chemistry Chapter 9 Stoichiometry

**LP Steffe**



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