

## Practicing the Mole - - Even Problems Answer Key

*Calculate the mass in grams of each of the following:*

2. 8.00 moles of aluminum

$$\frac{8.00 \text{ mol} \mid 26.98 \text{ g}}{1 \text{ mol}} = 216 \text{ g Al}$$

6. 7.00 moles of iodine (I<sub>2</sub>)

$$\frac{7.00 \text{ mol} \mid 253.8 \text{ g}}{1 \text{ mol}} = 1,780 \text{ g I}_2$$

4. 2.00 x 10<sup>2</sup> moles of chlorine (Cl<sub>2</sub>)

$$\frac{200. \text{ mol} \mid 70.90 \text{ g}}{1 \text{ mol}} = 14,200 \text{ g Cl}_2$$

8. 9.20 moles of iron

$$\frac{9.20 \text{ mol} \mid 55.85 \text{ g}}{1 \text{ mol}} = 514 \text{ g Fe}$$

*Calculate the number of moles of atoms in each of the following:*

10. 281 g of iron

$$\frac{281 \text{ g} \mid 1 \text{ mol}}{55.85 \text{ g}} = 5.03 \text{ mol Fe}$$

14. 93.0 g of phosphorus

$$\frac{93.0 \text{ g} \mid 1 \text{ mol}}{30.97 \text{ g}} = 3.00 \text{ mol P}$$

12. 72.0 g of magnesium

$$\frac{72.0 \text{ g} \mid 1 \text{ mol}}{24.31 \text{ g}} = 2.96 \text{ mol Mg}$$

16. 1035 g of lead

$$\frac{1035 \text{ g} \mid 1 \text{ mol}}{207.2 \text{ g}} = 4.995 \text{ mol Pb}$$

*Calculate the number of atoms in each of the following:*

18. 5.40 moles of iron

$$\frac{5.40 \text{ mol} \mid 6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$$

$$= 3.25 \times 10^{24} \text{ atoms Fe}$$

20. 10.0 moles of titanium

$$\frac{10.0 \text{ mol} \mid 6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$$

$$= 6.02 \times 10^{24} \text{ atoms Ti}$$

# Chemistry Inquiry Skill Practice Answers

**CO Houle**



## **Chemistry Inquiry Skill Practice Answers:**

**Holt Chemistry** Ralph Thomas Myers,2004     *Chemfile Skills Practice Experiments* Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2006     Professional Development of Chemistry Teachers Rachel Mamlok-Naaman,Ingo Eilks,George Bodner,Avi Hofstein,2018-05-09 Continuous professional development of chemistry teachers is essential for any effective chemistry teaching due to the evolving nature of the subject matter and its instructional techniques Professional development aims to keep chemistry teaching up to date and to make it more meaningful more educationally effective and better aligned to current requirements Presenting models and examples of professional development for chemistry teachers from pre service preparation through to continuous professional development the authors walk the reader through theory and practice The authors discuss factors which affect successful professional development such as workload availability and time constraints and consider how we maintain the life long learning of chemistry teachers With a solid grounding in the literature and drawing on many examples from the authors rich experiences this book enables researchers and educators to better understand teachers roles in effective chemistry education and the importance of their professional development

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