

COHESIVE DEVICES

ADDING

and
also
as well as
moreover
too
furthermore
in addition
additionally
plus
along with

SEQUENCING

first/firstly
second/secondly
third/thirdly
finally
next
before
after
then
subsequently

EMPHASISING

above all
in particular
especially
significantly
indeed
notably
chiefly
mainly
clearly
generally

CONTRASTING

however
whereas
instead of
nevertheless
alternatively
otherwise
unlike
on the other hand
conversely
yet

COMPARING

similarly
likewise
as with
just as
like
equally
not only ...
but also
in the same way

CAUSE and EFFECT

because
so
therefore
thus
consequently
hence
as a result
due to this

EXCEPTIONS

but
however
although
unless
if not
except
apart from
aside from
as long as

ILLUSTRATING

for example
such as
for instance
in the case of
as revealed by
illustrated by



Examples Of Cohesive Devices

**David Waugh, Claire Warner, Rosemary
Waugh**



Examples Of Cohesive Devices:

Unveiling the Magic of Words: A Report on "**Examples Of Cohesive Devices**"

In some sort of defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Examples Of Cohesive Devices**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room

temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable

measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less

exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol⁻¹

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eg 1 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = 57 \text{ kJ mol}^{-1}$ eg 2 $\text{HNO}_3(\text{aq}) + \text{KOH(aq)} \rightarrow \text{KNO}_3(\text{aq}) + \text{H}_2\text{O(l)}$ for which $\Delta H = 57 \text{ kJ mol}^{-1}$

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