

**CHEMFILE MINI-GUIDE TO PROBLEM SOLVING****2. PLAN**

- *What step is needed to calculate the concentration of the diluted solution?*

Apply the principle that  $\text{volume}_1 \times \text{molarity}_1 = \text{volume}_2 \times \text{molarity}_2$ .

$$M_1 V_1 = M_2 V_2 \xrightarrow[\text{substitute each of the known quantities for its symbol and calculate}]{\begin{array}{l} \text{rearrange the equation} \\ M_1 V_1 = M_2 V_2 \\ \text{algebraically to solve for } M_2 \end{array}} M_2 = \frac{M_1 V_1}{V_2}$$

$$M_2 = \frac{M_1 \times V_1}{V_2}$$

**3. COMPUTE**

Note: Even though molarity is moles per liter, you can use volumes in milliliters along with molarity whenever the units cancel.

$$M_2 = \frac{4.74 \text{ M} \times 50.00 \text{ mL}}{250.00 \text{ mL}} = 0.948 \text{ M}$$

**4. EVALUATE**

- *Are the units correct?*
- *Is the number of significant figures correct?*
- *Is the answer reasonable?*

Yes; molarity (mol/L) was required.  
 Yes; the number of significant figures is correct because the data had a minimum of three significant figures.  
 Yes; the computation is the same as  $4.74/5$ , which is a little less than 1.

**PRACTICE**

1. Complete the table below by calculating the missing value in each row.

|    | <b>Molarity of stock solution</b> | <b>Volume of stock solution</b> | <b>Molarity of dilute solution</b> | <b>Volume of dilute solution</b> |              |
|----|-----------------------------------|---------------------------------|------------------------------------|----------------------------------|--------------|
| a. | 0.500 M KBr                       | 20.00 mL                        | 7 M KBr                            | 100.00 mL                        | ans: 0.100 M |
| b. | 1.00 M LiOH                       | ? mL                            | 0.075 M LiOH                       | 500.00 mL                        | ans: 38 mL   |
| c. | 7 M HI                            | 5.00 mL                         | 0.0493 M HI                        | 100.00 mL                        | ans: 0.986 M |
| d. | 12.0 M HCl                        | 0.250 L                         | 1.8 M HCl                          | ? L                              | ans: 1.7 L   |
| e. | 7.44 M NH <sub>3</sub>            | ? mL                            | 0.093 M NH <sub>3</sub>            | 4.00 L                           | ans: 50. mL  |

# Chemfile Mini Guide To Problem Solving Answer Key

**L Towne**



## **Chemfile Mini Guide To Problem Solving Answer Key:**

Holt Chemistry File ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

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## **Table of Contents Chemfile Mini Guide To Problem Solving Answer Key**

1. Understanding the eBook Chemfile Mini Guide To Problem Solving Answer Key
  - The Rise of Digital Reading Chemfile Mini Guide To Problem Solving Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemfile Mini Guide To Problem Solving Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Chemfile Mini Guide To Problem Solving Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemfile Mini Guide To Problem Solving Answer Key
  - Personalized Recommendations
  - Chemfile Mini Guide To Problem Solving Answer Key User Reviews and Ratings
  - Chemfile Mini Guide To Problem Solving Answer Key and Bestseller Lists
5. Accessing Chemfile Mini Guide To Problem Solving Answer Key Free and Paid eBooks
  - Chemfile Mini Guide To Problem Solving Answer Key Public Domain eBooks
  - Chemfile Mini Guide To Problem Solving Answer Key eBook Subscription Services

- Chemfile Mini Guide To Problem Solving Answer Key Budget-Friendly Options
- 6. Navigating Chemfile Mini Guide To Problem Solving Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemfile Mini Guide To Problem Solving Answer Key Compatibility with Devices
  - Chemfile Mini Guide To Problem Solving Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemfile Mini Guide To Problem Solving Answer Key
  - Highlighting and Note-Taking Chemfile Mini Guide To Problem Solving Answer Key
  - Interactive Elements Chemfile Mini Guide To Problem Solving Answer Key
- 8. Staying Engaged with Chemfile Mini Guide To Problem Solving Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemfile Mini Guide To Problem Solving Answer Key
- 9. Balancing eBooks and Physical Books Chemfile Mini Guide To Problem Solving Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemfile Mini Guide To Problem Solving Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemfile Mini Guide To Problem Solving Answer Key
  - Setting Reading Goals Chemfile Mini Guide To Problem Solving Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemfile Mini Guide To Problem Solving Answer Key
  - Fact-Checking eBook Content of Chemfile Mini Guide To Problem Solving Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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