

# QUALITATIVE ANALYSIS

Experiment	Observation	Inference
<b><u>Preliminary Test</u></b>		
1. Noted the colour of the given salt	1. Blue or Bluish Green 2. Greenish 3. Light Green 4. Dark Brown 5. Light Pink 6. Colourless	$\text{Cu}^{2+}$ or $\text{Ni}^{2+}$ $\text{Ni}^{2+}$ $\text{Fe}^{2+}$ $\text{Fe}^{3+}$ $\text{Mn}^{2+}$ Absence of $\text{Cu}^{+2}$ , $\text{Ni}^{2+}$ , $\text{Fe}^{2+}$ , $\text{Fe}^{3+}$ , $\text{Mn}^{2+}$
2. Noted the solubility	1. Soluble in water 2. Soluble in dilute HCl	$\text{NH}_4^+$ may be present Absence of group I cations
3. Dry Heating Test: Heated a pinch of the salt in a dry Test tube	1. Colour, odourless gas 2. Colourless gas with rotten egg smell 3. Colourless gas with vinegar smell 4. Colourless gas with ammonia smell 5. Reddish brown gas 6. Greenish yellow gas 7. Violet gas	$\text{CO}_3^{2-}$ may be present $\text{S}^{2-}$ may be present $\text{CH}_3\text{COO}^-$ may be present $\text{NH}_4^+$ may be present $\text{NO}_3^-$ or $\text{Br}^-$ may be present $\text{Cl}^-$ may be present $\text{I}^-$ may be present
<b><u>Identification of acid radical (Anion)</u></b>		
1. Dilute Acid Test: Treated a pinch of the substance With dilute $\text{H}_2\text{SO}_4$	1. Colourless, odourless gas with brisk effervescence which turned lime water milky. 2. Vinegar Smell 3. No characteristic Reaction	$\text{CO}_3^{2-}$ present $\text{CO}_3^{2-} + 2\text{H}^+ \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$ $\text{CH}_3\text{COO}^-$ Present $\text{CH}_3\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{COOH}$ Absence of $\text{CO}_3^{2-}$ , $\text{CH}_3\text{COO}^-$
2. Concentrated Acid Test: Treated a pinch of the substance With concentrated $\text{H}_2\text{SO}_4$	1. Colourless gas with pungent Smell which gave dense white fumes when a glass rod dipped in $\text{NH}_4\text{OH}$ was brought near the mouth of the test tube 2. Reddish brown vapours with Pungent smell. Solution turns red 3. Deep violet vapours with pungent smell. Sides of test tube	$\text{Cl}^-$ present $\text{Cl}^- + \text{H}^+ \rightarrow \text{HCl}$ $\text{HCl} + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$ $\text{Br}^-$ present $\text{Br}^- + \text{H}^+ \rightarrow \text{HBr}$ $2\text{HBr} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Br}_2 + \text{SO}_2$ $\text{I}^-$ present $\text{I}^- + \text{H}^+ \rightarrow \text{HI}$ $2\text{HI} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{I}_2 + \text{SO}_2$

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