

## CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron III sulfate, the roman numeral tells us **C**
  - How many atoms of Fe are in one formula unit
  - How many sulfate ions can be attached to the iron atom
  - The charge on each Fe ion
  - The total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to **c**
  - Change the number of moles represented in the formula
  - Change the charges on the other ions in the compound
  - Change the formula so that it no longer represents the compound
  - Have no effect on the formula
- Derive a generalization for when an acid name will end in the suffix **-ic** or **-ous**  
**If ending is -ic, the anion ended in -ate; if ending is -ous, the anion ended in -ite**
- Derive a generalization for when an acid name will begin with the prefix **hydro-**.  
**If prefix is hydro-, the anion ended in -ide. This is most likely a monatomic ion.**
- Fill in the blanks in the table below

Compound Name	Formula
Aluminum sulfide	$Al_2S_3$
Aluminum sulfate	$Al_2(SO_4)_3$
Lead II chloride	$PbCl_2$
Ammonium phosphite	$(NH_4)_2PO_3$
Hydroiodic acid	$HCl$
Dinitrogen pentoxide	$N_2O_5$
Copper II carbonate	$CuCO_3$
Sodium sulfide	$Na_2S$
Ammonium phosphate	$(NH_4)_3PO_4$
Tin IV sulfite	$Sn(SO_3)_2$
Nitrous acid	$HNO_2$
Magnesium perchlorate	$Mg(ClO_4)_2$
Iron II nitrate	$Fe(NO_3)_2$
Iron III nitrate	$Fe(NO_3)_3$
Cobalt II oxide	$CoO$
Sulfur trioxide	$SO_3$
Boron pentafluoride	$BF_5$
Tetraphosphorus octasulfide	$P_4S_8$
Tetraiodide nonoxide	$I_4O_9$

- Name the following compound and identify as molecular, ionic or acid
  - $Ca(HCO_3)_2$  **calcium bicarbonate - I**
  - $Fe(ClO_2)_3$  **Iron III chlorite - I**
  - $Fe(ClO_4)_2$  **Iron II chlorate - I**
  - $N_2O_5$  **dinitrogen pentoxide - M**

# Chapter 9 Chemical Names Formulas Answers Key

**J Ma**



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