

Notes Ch. 4 and 25: Atomic Structure and Nuclear Chemistry
History and Structure of the Nuclear Atom

The Atom

- smallest particle of an element that retains all properties of the element

I. Early Models of the Atom

A. Democritus (460 B.C. – 370 B.C.)

- first to suggest the existence of _____
- **believed** atoms were indivisible and indestructible

B. Dalton

- Atoms of the same element are _____
- Each element is unique
- Atoms combine in fixed ratios to form compounds

C. Thomson

- Discovered _____
- Suggested atom looked like plum pudding (or a blueberry muffin or chocolate chip cookie) with electrons evenly distributed throughout positive sphere

D. Rutherford

- Conducted gold foil experiment
- Proposed atom is mostly empty space
- Concluded all positive charge and mass is concentrated in small region called _____

E. Bohr

- depicts the atom as a small, positively charged nucleus
- surrounded by electrons that travel in circular _____ around the nucleus

F. Schrödinger

- model allowed the electron to occupy three-dimensional space like an electron " _____ "
- Used a mathematical equation to show a model of electrons as waves

II. More about the Atom

A. Size

- Teeny tiny
- observable with instruments such as a scanning tunneling _____

B. Parts (called Subatomic Particles)

1. Protons: p^+

- charge = positive +1
- mass = 1

2. Electrons: e^-

- charge = negative -1
- mass ~ 0

3. Neutrons: n^0

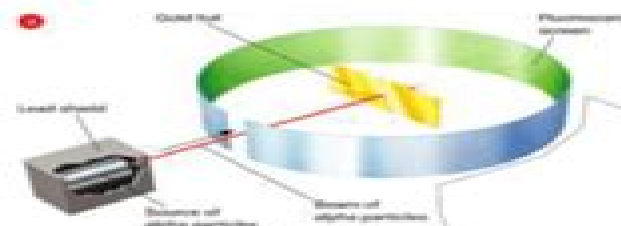
- charge = zero
- mass = 1

C. Location of Parts

- **Nucleus:** center of atom
 - contains protons and neutrons
 - has a positive charge
 - contains almost ALL of the _____ of the atom
- **Outside nucleus:**
 - contains _____
 - has a negative charge
 - occupies almost ALL of the volume of the atom

D. Atomic Charge

- _____ are NEUTRAL
- # of protons = # of electrons
(ALWAYS if you're talking about an ATOM)



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Table of Contents Chapter 25 Nuclear Chemistry Guided Reading Answers

1. Understanding the eBook Chapter 25 Nuclear Chemistry Guided Reading Answers
 - The Rise of Digital Reading Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 25 Nuclear Chemistry Guided Reading Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Personalized Recommendations
 - Chapter 25 Nuclear Chemistry Guided Reading Answers User Reviews and Ratings
 - Chapter 25 Nuclear Chemistry Guided Reading Answers and Bestseller Lists
5. Accessing Chapter 25 Nuclear Chemistry Guided Reading Answers Free and Paid eBooks

- Chapter 25 Nuclear Chemistry Guided Reading Answers Public Domain eBooks
 - Chapter 25 Nuclear Chemistry Guided Reading Answers eBook Subscription Services
 - Chapter 25 Nuclear Chemistry Guided Reading Answers Budget-Friendly Options
6. Navigating Chapter 25 Nuclear Chemistry Guided Reading Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 25 Nuclear Chemistry Guided Reading Answers Compatibility with Devices
 - Chapter 25 Nuclear Chemistry Guided Reading Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Highlighting and Note-Taking Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Interactive Elements Chapter 25 Nuclear Chemistry Guided Reading Answers
 8. Staying Engaged with Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 25 Nuclear Chemistry Guided Reading Answers
 9. Balancing eBooks and Physical Books Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 25 Nuclear Chemistry Guided Reading Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Setting Reading Goals Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Fact-Checking eBook Content of Chapter 25 Nuclear Chemistry Guided Reading Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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