

CHAPTER 1: STOICHIOMETRY

- One atom of sodium is
 - 23g
 - $\frac{1}{23}$ moles
 - Positively charged
 - $\frac{1}{23}$ moles
- Diamerck brown is a dye. Its molar mass is 228.3 g/mol. When the dye was analyzed by a scientist, it was found that it contains 30.65% nitrogen. How many nitrogen atoms are there in each Diamerck brown molecule?
 - 0
 - 5
 - 4
 - 3
- The mass of 2 moles of sodium hydroxide will be
 - 2g
 - 20g
 - 40g
 - 80g
- Consider the following reaction:

$$2\text{Al}_{(s)} + 3\text{Cl}_{2(g)} \rightarrow 2\text{AlCl}_{3(s)}$$

$23 + 16 + 1 = 40$
 $2 \times 40 = 80\text{g}$

The following figure represents a mixture of Al and Cl_2 .



- Consider that each symbol is representing one mole of the substance. Which of the reactants will be consumed earlier and how many moles of the product will be formed?
- Al and 4 mol of AlCl_3
 - Cl_2 and 4 mol of AlCl_3
 - Cl_2 and 6 mol of AlCl_3
 - both reactants will be consumed at the same time and 4 mol of AlCl_3
 - moles of H_2O contain
 - 100 moles of bonds
 - 100 moles of electrons

- 30 atoms
 - 40 moles of hydrogen bonds
- Ozone (O_3) filters the cosmic rays of sunlight. How many oxygen atoms are there in 0.2 mole of ozone (O_3 , molar mass 48.0 g/mol)?
 - 6.02×10^{23}
 - 6.02×10^{22}
 - 3.61×10^{23}
 - 6×10^{23}
 - On heating, the hydrated salt $\text{CaSO}_4 \cdot x\text{H}_2\text{O}$ loses its water of crystallization. In an experiment it was found that when 0.2 mol of the hydrated salt was heated 0.10 mol of water was lost. What is the molecular formula of the hydrated compound?
 - CaSO_4
 - $\text{CaSO}_4 \cdot 9\text{H}_2\text{O}$
 - $\text{CaSO}_4 \cdot \text{H}_2\text{O}$
 - $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$
 - Zirconium chloride reacts with magnesium to produce Zirconium

$$\text{ZrCl}_4 + 2\text{Mg} \rightarrow \text{Zr} + 2\text{MgCl}_2$$

$3 \text{ mole} \qquad \qquad \qquad 2 \text{ mole}$

How many moles of Zr would be produced if 0.2 mol of Mg metal is used in the reaction?

 - 0.1
 - 0.01
 - 0.20
 - 0.05
 - Tobacco contains a poisonous alkaloid called Nicotine. The molecular formula of nicotine is $\text{C}_{10}\text{H}_{14}\text{N}_2$ (molar mass = 162.23 g). The mass percentage of nitrogen in nicotine, is?
 - 4.32%
 - 8.23%
 - 12.4%
 - 17.3%
 - How much calcium is present in 10g of CaCO_3 ?
 - 1g
 - 2g
 - 3g
 - 4g
 - How many carbonate ions are there in 50g of CaCO_3 ?

$\text{N}_2 = 28\text{g}$
 $1 \text{ mol} \quad \frac{162.23}{28} \times 100 = 17.3\%$

$0.2 \text{ mol Mg} / 2 \text{ mol Mg} \times 2 \text{ mol Zr} = 0.2 \text{ mol Zr}$

$100 = 17.3\%$

- $50 = 5.022 \times 10^{23}$
 - $100 = 5.022 \times 10^{23}$
 - $50 = 5.022 \times 10^{23}$
 - $50 = 5.022 \times 10^{23}$
- Stoichiometric calculations occur when:
 - The reaction is irreversible
 - There is no side reaction
 - At least one of the reactants is completely consumed
 - All of the above
 - How many molecules of $\text{K}_2\text{Cr}_2\text{O}_7$ are formed from the decomposition of potassium chlorate by reaction?
 - $3 \times 6.022 \times 10^{23}$
 - 3×10^{23}
 - $3 \times 10^{23} \times 6.022 \times 10^{23}$
 - 6
 - Which of the following substances?
 - 4g helium
 - 6.022×10^{23} mol
 - 40g calcium
 - 6.022×10^{23}
 - Urea (NH_2CO) and urea are present in the mixture.
 - 14g
 - 56g
 - 28g
 - 42g
 - 8g of methane
 - 4 hydrogens
 - 0.5 mol of oxygen
 - 6.022×10^{23}
 - $4 \times 6.022 \times 10^{23}$
 - A beaker has 10g of oxygen
 - 10
 - 18



Chapter 9 Stoichiometry Multiple Choice Answers

Ying Liu



Chapter 9 Stoichiometry Multiple Choice Answers:

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