

## Chapter 6 – Study Guide

Test Date: Friday, December 19, 2014

Use the following to study:

Textbook – Ch. 6

Ch. 6 homework packet

### Chemical Bonding:

- **Four types of bonds:**

- *Nonpolar Covalent Bond* ( $\Delta EN$  (difference in electronegativity) =  $< 0.3$ )
  - EQUAL SHARING of electrons between the 2 bonding nonmetal atoms
- *Polar Covalent Bond* ( $\Delta EN = 0.4-1.7$ )
  - UNEQUAL SHARING of electrons between the 2 bonding nonmetal atoms (because EN is much higher in one atom, the electrons are pulled closer to the atom in the bond with the higher EN value – like an uneven tug of war)
- *Ionic Bond* ( $\Delta EN = > 1.7$ )
  - Difference in the EN values for the 2 bonding atoms (metal & nonmetal) is so large that the atom with the lower EN value ends up completely GIVING UP/TRANSFERRING ELECTRONS to the atom with the greater EN value
  - Ionic bond results due to the attraction between the POSITIVE cation (formed by losing electrons) and the NEGATIVE anion (formed by gaining electrons)
- *Metallic Bond*
  - Bonding between two metallic atoms which occurs due to the attraction between a "sea of electrons" and the nuclei of bonded metal atoms

\* Must memorize/know the difference in electronegativity ranges for each type of bond

\*\* Note: you should still know the definition of electronegativity (EN)

- **Properties of the Four Types of Chemical Bonds**

- See the table at the bottom of Packet pg.5
- You MUST be able to compare all four examples – NaCl (ionic), Sucrose (Polar), Stearic Acid/Wax (nonpolar), and Aluminum & Mercury (metallic) – in terms of melting point, conductivity of solid, solubility in water, conductivity in water (when applicable), solubility in hexane, hardness (ease with which it is scratched), and brittleness

- **Why & How do chemical bonds form?**

- **Why** → because all atoms want to be at their **LOWEST POTENTIAL ENERGY**; for nearly all atoms (except the noble gases) this happens only when an atom chemically bonds to other atoms!!!
- **How** → **ATTRACTIVE** and **REPULSIVE** forces
  - *Attractive forces* – positively charged nucleus of one atom is attracted to electrons in another atom nearby (because they have opposite charges)
  - *Repulsive forces* – 2 nuclei or 2 sets of electron clouds from nearby atoms repel one another (because they have the same charge)
- **Bond Length**
  - ideal distance between 2 atoms that gives the lowest potential energy of the atoms by balancing out attractive & repulsive forces)
- **Bond Energy**
  - Energy required to break apart a chemical bond (the higher the bond energy, the stronger the bond)

# Chemistry Chapter 6 Study Guide Answers

**John McMurry, Susan McMurry**



## **Chemistry Chapter 6 Study Guide Answers:**

**The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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Susan McMurry, 1996 This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker more realistic view of Botswana Sue Baker Publishing News

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**Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry** McMurry, Susan McMurry, 1990

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**Study Guide and Solutions Book** Harold Hart, 1983

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*Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual* T. W. Graham Solomons, 1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

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Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book *Chemistry, Student Study Guide* John A. Olmsted,Gregory M. Williams,2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that s real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You ll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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