



6: Equilibrium

Chemistry Worksheet

Standard: 11th Science

Subject: Chemistry

Student Name: _____

Q1. Which of the following is not a buffer solution?

1. $\text{CH}_3\text{COOH} + \text{CH}_3\text{COONa}$
2. $\text{H}_2\text{BO}_3 + \text{Na}_2\text{BO}_3$
3. $\text{HClO}_4 + \text{NaClO}_4$
4. $\text{NH}_4\text{OH} + (\text{NH}_4)_2\text{SO}_4$

Q2. 0.10M CH_3COOH is 1.34% ionised, calculate its K_a :

1. 1.8×10^{-1}
2. 1.8×10^{-4}
3. 5×10^{-4}
4. 4×10^{-5}

Q3. Fear or excitement generally causes one to breathe rapidly and it results in the decrease of concentration of CO_2 in blood.

In what way, it will change the pH of blood?

1. pH will increase
2. pH will decrease
3. No change
4. pH will be 7

Q4. Pure ammonia is placed in a vessel at a temperature where its dissociation constant (α) is appreciable. At equilibrium :

1. K_p does not change significantly with pressure.
2. α does not change with pressure.
3. The concentration of NH_3 does not change with pressure.
4. The concentration of hydrogen is less than that of nitrogen.

Q5. Which of the following salts will give highest pH in water?

1. KCl
2. NaCl
3. Na_2CO_3
4. CuSO_4

Q6. "An acid is a substance that is capable of donating a hydrogen ion H^+ and bases are substances capable of accepting a



Chemistry Chemical Equilibrium Test Answers

R Sandford



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