

Atomic structure

Label the following diagram with the **correct** proton, neutron, electron and nucleus.



Carbon atom

2. Now complete the table below, identifying the charges, mass and place found for each of the sub-atomic particles.

Sub-atomic particle	Place found	Mass	Charge
proton			
neutron			
electron			

3. Why are atoms neutral overall?

.....

.....

Atomic structure 2

Use the information in the periodic table to help you complete the table below:

Element	Protons	Neutrons	Electrons
Carbon			
Sodium			
Magnesium			
Isotope			
Bromine			
Hydrogen			
Aluminium			
Argon			

3. What is an isotope?

.....

.....

Calculating Relative Atomic Mass (A_r)

Calculate the relative atomic mass of the following elements, given their mass and abundance.

a) Bromine, consisting of 50% Br-79 and 50% Br-81

.....

b) Copper, consisting of 69.09% Cu-63 and 30.91% Cu-65

.....

c) Silver, consisting of 97.54% Ag-107 and 2.46% Ag-109

.....

The Periodic Table

1. What element is found in:

- a) Group 2 Period 4
- b) Group 1 Period 1
- c) Group 7 Period 2
- d) Group 8 Period 3

2. What group and period are the following elements found in?

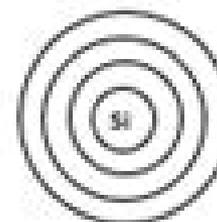
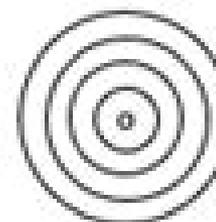
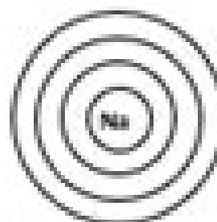
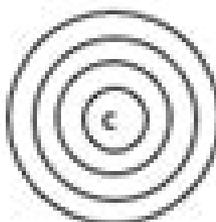
- a) Lithium
- b) Phosphorus
- c) Iron
- d) Rubidium

Electron structure

1. How many electrons can each shell hold?

- a) 1st shell:
- b) 2nd shell:
- c) 3rd shell:
- d) 4th shell:

2. Complete the electron structure diagrams for the following elements by drawing electrons on the shells. Under each diagram, write the electron configuration, e.g. 2, 8, 1



The Periodic Table

- Shade the metals in blue
- Shade the non-metals in red
- Shade the metalloids in green
- What is special about the properties of metalloids?

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**Thirumalaisamy P.
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