

CHAPTER 3 CHAPTER ASSESSMENT

MATTER AND CHANGE

Reviewing Vocabulary

Match the definition in Column A with the term in Column B.

Column A

- _____ 1. The outermost electrons of an atom
- _____ 2. The center of an atom
- _____ 3. Hot, highly ionized, electrically conducting gas
- _____ 4. The smallest particle of an element that retains that element's characteristics
- _____ 5. The change from a gas to a liquid
- _____ 6. A homogeneous mixture
- _____ 7. The change of one or more substances into other substances
- _____ 8. The attractive force between two ions of opposite charge
- _____ 9. A tiny particle with a negative electrical charge
- _____ 10. The average of the mass numbers of an element's isotopes
- _____ 11. The spontaneous process through which unstable nuclei emit radiation

Column B

- a. atom
- b. atomic mass
- c. chemical reaction
- d. condensation
- e. electron
- f. ionic bond
- g. solution
- h. nucleus
- i. plasma
- j. valence electrons
- k. radioactive decay

In the space at the left, write the word or phrase in parentheses that makes the statement correct.

- _____ 12. The combined number of protons and neutrons is the (atomic number, mass number).
- _____ 13. An (energy level, isotope) represents the area in an atom where an electron is most likely to be found.
- _____ 14. (An ionic bond, A covalent bond) is the attraction of two atoms for a shared pair of electrons that holds the atoms together.
- _____ 15. An atom that gains or loses an electron is a charged particle called (an ion, a molecule).
- _____ 16. (An element, A compound) is a substance that cannot be broken down into simpler substances.
- _____ 17. A tiny particle that has mass and a positive electrical charge is a (neutron, proton).
- _____ 18. (Acids, Bases) are solutions characterized by the formation of hydroxide ions (OH^-).
- _____ 19. The process of changing from a liquid into a gas is called (evaporation, sublimation).

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**Edgar Fahs Smith, Harry Frederick
Keller**



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