

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## States of Matter

## Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

- The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
- The kinetic molecular theory makes the following assumptions:
  - In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
  - Between gas particles and the walls, there are no significant attractive or repulsive forces between gas particles.
  - Gas particles are in constant and random motion.
  - The collisions between gas particles are elastic, that is, no kinetic energy is lost.
- The kinetic energy of a particle is represented by the equation  $KE = \frac{1}{2}mv^2$ .
- Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- |              |  |
|--------------|--|
| <u>True</u>  | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas.                              |
| <u>True</u>  | 2. The random motion of gas particles causes a gas to expand until it fills its container.                                       |
| <u>False</u> | 3. The density of a gas decreases as it is compressed.   |
| <u>True</u>  | 4. A gas can flow into a space occupied by another gas.  |
| <u>True</u>  | 5. The diffusion of a gas is caused by the random motion of the particles of the gas.  |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles.   |
| <u>True</u>  | 7. During effusion, a gas escapes through a tiny opening into a vacuum.  |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

Study Guide for Content Mastery

Chemistry: Matter and Change • Chapter 13 79

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## Section 13.1 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

- Pressure is defined as force per unit
  - area
  - time
  - mass
  - volume
- What is an instrument designed to measure atmospheric pressure?
  - barometer
  - manometer
  - physoconimeter
  - densimeter
- The height of the liquid in a barometer is affected by all of the following EXCEPT the
  - altitude
  - density of the liquid in the column
  - atmospheric pressure
  - diameter of the column tube.
- The pressure of the gas in a manometer is directly related to which of the following
  - height of the mercury column in the closed-end arm
  - height of the mercury column in the open-end arm
  - $h + 76$
  - $h - 76$
- One atmosphere is equal to a pressure of
  - 76 mm Hg
  - 101.3 kPa
  - 147 psi
  - 760 mm
- The partial pressure of a gas depends on all of the following EXCEPT the
  - concentration of the gas
  - size of the container
  - density of the gas
  - temperature of the gas
- The pressure of a sample of air in a manometer is 101.3 kPa. What is the partial pressure of nitrogen ( $N_2$ ) in the sample if the combined partial pressures of the other gases is 22.4 kPa?
  - 42.4 kPa
  - 78.9 kPa
  - 101.3 kPa
  - 123.7 kPa

Use the figure to answer the following questions.

- What instrument is illustrated in the figure? barometer
- Who invented this instrument? Evangelista Torricelli
- What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
- What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



80 Chemistry: Matter and Change • Chapter 13

Study Guide for Content Mastery

# Chemistry Matter And Change Chapter 13 Study Answers

**Lei Shi**



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## **Table of Contents Chemistry Matter And Change Chapter 13 Study Answers**

1. Understanding the eBook Chemistry Matter And Change Chapter 13 Study Answers
  - The Rise of Digital Reading Chemistry Matter And Change Chapter 13 Study Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Chapter 13 Study Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Matter And Change Chapter 13 Study Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Chapter 13 Study Answers
  - Personalized Recommendations
  - Chemistry Matter And Change Chapter 13 Study Answers User Reviews and Ratings

- Chemistry Matter And Change Chapter 13 Study Answers and Bestseller Lists
- 5. Accessing Chemistry Matter And Change Chapter 13 Study Answers Free and Paid eBooks
  - Chemistry Matter And Change Chapter 13 Study Answers Public Domain eBooks
  - Chemistry Matter And Change Chapter 13 Study Answers eBook Subscription Services
  - Chemistry Matter And Change Chapter 13 Study Answers Budget-Friendly Options
- 6. Navigating Chemistry Matter And Change Chapter 13 Study Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Matter And Change Chapter 13 Study Answers Compatibility with Devices
  - Chemistry Matter And Change Chapter 13 Study Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Matter And Change Chapter 13 Study Answers
  - Highlighting and Note-Taking Chemistry Matter And Change Chapter 13 Study Answers
  - Interactive Elements Chemistry Matter And Change Chapter 13 Study Answers
- 8. Staying Engaged with Chemistry Matter And Change Chapter 13 Study Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Matter And Change Chapter 13 Study Answers
- 9. Balancing eBooks and Physical Books Chemistry Matter And Change Chapter 13 Study Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Matter And Change Chapter 13 Study Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Matter And Change Chapter 13 Study Answers
  - Setting Reading Goals Chemistry Matter And Change Chapter 13 Study Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Matter And Change Chapter 13 Study Answers
  - Fact-Checking eBook Content of Chemistry Matter And Change Chapter 13 Study Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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