

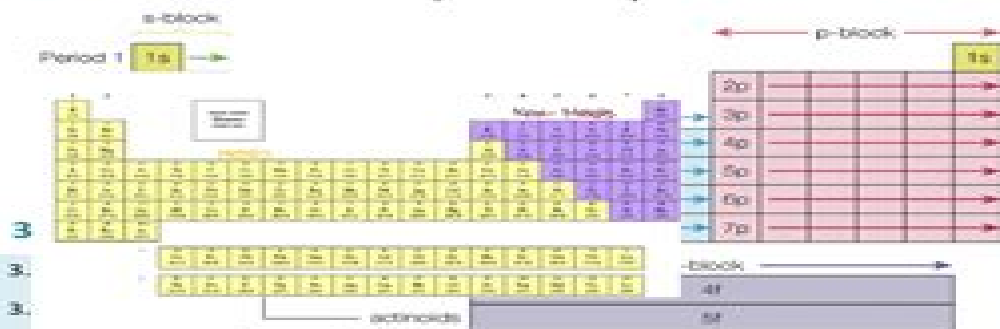
# Topic 3: Periodicity

## 3.1 Periodic Table

- 3.1.1 The periodic table is arranged into four blocks associated with the four sub-levels – s, p, d, and f
- 3.1.2 The periodic table consists of groups (vertical columns) and periods (horizontal rows)
- 3.1.3 The period number (n) is the outer energy level that is occupied by electrons
- 3.1.4 The number of the principal energy level and the number of the valence electrons in an atom can be deduced from its position on the periodic table
- 3.1.5 The periodic table shows the position of metals, non-metals and metalloids
- 3.1.6 Deduction of the electron configuration of an atom from the element's position on the periodic table, and vice versa

### Periodic Table Arrangement

- The periodic table is arranged into 4 blocks associated with their last orbital: s, p, d and f
- Elements in the periodic table are arranged by increasing atomic number
- **Group:** A vertical column of elements
  - Group number is the same as the number of valence electrons
- **Period:** A horizontal row of elements
  - Period number is the same number of shells in the atom
  - All shells except the outer shell will be full
- There is a division between metals and non-metals. Metals are on the left and non-metals are on the right
- The farther sub-level occupied is found by the location of the element (which block it is in):



- The location of the following families of the periodic table must be known: alkali metals, alkali earth metals, halogens, noble gases, transition metals, lanthanides and actinoids

- 3.2.1 Atomic radius, ionic radius
- 3.2.2 Electronegativity, ionization energy
- 3.2.3 Oxides change from basic through amphoteric to acidic across a period
- 3.2.4 Prediction and explanation of the metallic and non-metallic behavior of an element based on its position in the periodic table
- 3.2.5

**Atomic Radius** – The distance from the nucleus to the outermost electron

**Ionic Radius** – The distance from the nucleus to the outermost electrons in an ion

**First ionization energy** – The energy required to remove one mole of electrons from one mole of gaseous atoms. This is shown by:

**Electron affinity** – The energy released when one mole of an electron is added to one mole gaseous atoms. This is shown by:

**Electronegativity** – A measure of the attraction an atom has for a shared pair of electrons in a covalent bond

# Chemistry Topic 3 Periodic Table Outline Answers

**Michael Seilmaier**



## **Chemistry Topic 3 Periodic Table Outline Answers:**

## **Chemistry Topic 3 Periodic Table Outline Answers** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the energy of words has are more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such may be the essence of the book **Chemistry Topic 3 Periodic Table Outline Answers**, a literary masterpiece that delves deep in to the significance of words and their affect our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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### **Table of Contents Chemistry Topic 3 Periodic Table Outline Answers**

1. Understanding the eBook Chemistry Topic 3 Periodic Table Outline Answers
  - The Rise of Digital Reading Chemistry Topic 3 Periodic Table Outline Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Topic 3 Periodic Table Outline Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Topic 3 Periodic Table Outline Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Topic 3 Periodic Table Outline Answers
  - Personalized Recommendations
  - Chemistry Topic 3 Periodic Table Outline Answers User Reviews and Ratings

- Chemistry Topic 3 Periodic Table Outline Answers and Bestseller Lists
- 5. Accessing Chemistry Topic 3 Periodic Table Outline Answers Free and Paid eBooks
  - Chemistry Topic 3 Periodic Table Outline Answers Public Domain eBooks
  - Chemistry Topic 3 Periodic Table Outline Answers eBook Subscription Services
  - Chemistry Topic 3 Periodic Table Outline Answers Budget-Friendly Options
- 6. Navigating Chemistry Topic 3 Periodic Table Outline Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Topic 3 Periodic Table Outline Answers Compatibility with Devices
  - Chemistry Topic 3 Periodic Table Outline Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Topic 3 Periodic Table Outline Answers
  - Highlighting and Note-Taking Chemistry Topic 3 Periodic Table Outline Answers
  - Interactive Elements Chemistry Topic 3 Periodic Table Outline Answers
- 8. Staying Engaged with Chemistry Topic 3 Periodic Table Outline Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Topic 3 Periodic Table Outline Answers
- 9. Balancing eBooks and Physical Books Chemistry Topic 3 Periodic Table Outline Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Topic 3 Periodic Table Outline Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Topic 3 Periodic Table Outline Answers
  - Setting Reading Goals Chemistry Topic 3 Periodic Table Outline Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Topic 3 Periodic Table Outline Answers
  - Fact-Checking eBook Content of Chemistry Topic 3 Periodic Table Outline Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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