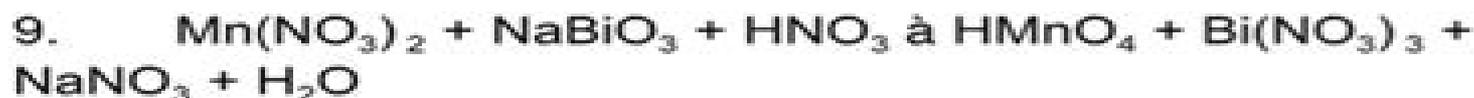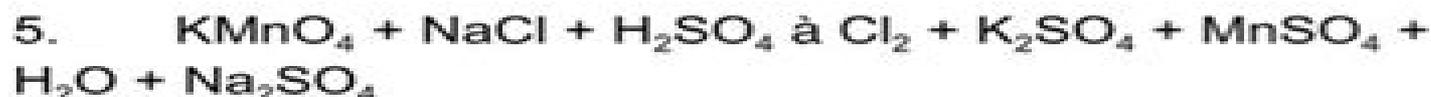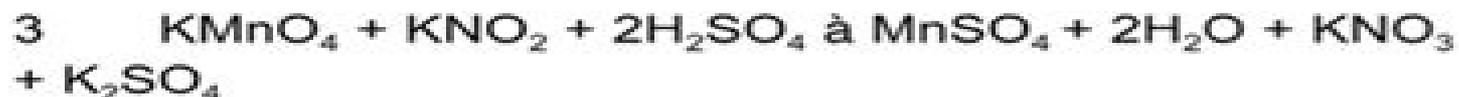# Oxidation-Reduction Worksheet

For each reaction below, identify the atom oxidized, the atom reduced, the oxidizing agent, the reducing agent, the oxidation half reaction, the reduction half reaction, and then balance the equation by the method of oxidation-reduction showing all electrons transfers.

1. $Mg + 2HCl → MgCl_2 + H_2$

2. $2Fe + 3V_2O_3 → Fe_2O_3 + 6VO$

3. $KMnO_4 + KNO_2 + 2H_2SO_4 → MnSO_4 + 2H_2O + KNO_3 + K_2SO_4$

4. $K_2Cr_2O_7 + SnCl_2 + HCl → CrCl_3 + SnCl_4 + KCl + H_2O$

5. $KMnO_4 + NaCl + H_2SO_4 → Cl_2 + K_2SO_4 + MnSO_4 + H_2O + Na_2SO_4$

6. $K_2Cr_2O_7 + H_2O + S → SO_2 + KOH + Cr_2O_3$

7. $KClO_3 + C_{12}H_{22}O_{11} → KCl + H_2O + CO_2$

8. $H_2C_2O_4 + K_2MnO_4 → CO_2 + K_2O + Mn_2O_3 + H_2O$

9. $Mn(NO_3)_2 + NaBiO_3 + HNO_3 → HMnO_4 + Bi(NO_3)_3 + NaNO_3 + H_2O$

10. $H_2C_2O_4 + KMnO_4 → CO_2 + K_2O + Mn_2O_3 + H_2O$

Oxidation And Reduction Practice With Answers

A Loxley

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