

## AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of  $\text{O}_2$ ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{3 \text{ O}_2} \times \frac{2 \text{ SO}_2}{2 \text{ SO}_2} \times \frac{3 \text{ S}}{\text{SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of  $\text{O}_2(g)$  are needed to produce 4.6 moles of  $\text{Fe}(s)$ ?A) 2.3 mole  $\text{O}_2$ B) 3.5 mole  $\text{O}_2$ C) 4.6 mole  $\text{O}_2$ D) 6.9 mole  $\text{O}_2$ E) 13.8 mole  $\text{O}_2$ 

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ CO}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{326 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} =$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{3 \text{ H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ H}_2\text{SO}_4} =$$

B

- 4.
- $\text{SO}_2$
- reacts with
- $\text{H}_2\text{S}$
- as follows:

When 7.50 g of  $\text{H}_2\text{S}$  reacts with 12.75 g of  $\text{SO}_2$ , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.

C) 0.0216 moles of  $\text{H}_2\text{S}$  remain.D) 1.13 g of  $\text{H}_2\text{S}$  remain.E)  $\text{SO}_2$  is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.03 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{2 \text{ H}_2\text{S}} \times \frac{3 \text{ S}}{1 \text{ mol H}_2\text{S}} = 10.89$$

$$\frac{12.75 \text{ g SO}_2}{64 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ SO}_2} \times \frac{3 \text{ S}}{1 \text{ mol S}} = 19.13$$

# Chapter 4 Chemistry Test Answers

**R Pring**



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## **Table of Contents Chapter 4 Chemistry Test Answers**

1. Understanding the eBook Chapter 4 Chemistry Test Answers
  - The Rise of Digital Reading Chapter 4 Chemistry Test Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Chemistry Test Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 4 Chemistry Test Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Chemistry Test Answers
  - Personalized Recommendations
  - Chapter 4 Chemistry Test Answers User Reviews and Ratings
  - Chapter 4 Chemistry Test Answers and Bestseller Lists
5. Accessing Chapter 4 Chemistry Test Answers Free and Paid eBooks
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  - Chapter 4 Chemistry Test Answers eBook Subscription Services

- Chapter 4 Chemistry Test Answers Budget-Friendly Options
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  - Chapter 4 Chemistry Test Answers Compatibility with Devices
  - Chapter 4 Chemistry Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 4 Chemistry Test Answers
  - Highlighting and Note-Taking Chapter 4 Chemistry Test Answers
  - Interactive Elements Chapter 4 Chemistry Test Answers
- 8. Staying Engaged with Chapter 4 Chemistry Test Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 4 Chemistry Test Answers
- 9. Balancing eBooks and Physical Books Chapter 4 Chemistry Test Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 4 Chemistry Test Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 4 Chemistry Test Answers
  - Setting Reading Goals Chapter 4 Chemistry Test Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 4 Chemistry Test Answers
  - Fact-Checking eBook Content of Chapter 4 Chemistry Test Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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