

**REVIEW SHEET FOR 7th GRADE SCIENCE FINAL EXAM
2015**



LIFE SCIENCE - *From Bacteria to Plants*

1. Levels of organization - cell to organism
2. Activities and needs of living things
3. Unicellular, multicellular
4. Classification of living things - kingdom to species
5. Scientific name - Genus species
6. Six kingdoms of all living things
7. Virus, bacteriophage
8. Structure of a virus, how viruses multiply
9. Prokaryote, eukaryote
10. Autotroph, heterotroph
11. Bacteria, structure of a bacterial cell, reproduction in bacteria
12. Flagella
13. How bacteria are helpful and harmful
14. Infectious diseases
15. Protists, animal like protists (protozoans) and plant-like protists (algae)
16. Structure of amoeba, paramecium and euglena
17. Eutrophication, red tide
18. Fungi - structures and activities
19. Methods of reproduction - budding, binary fission, sexual, asexual
20. How fungi are helpful and harmful
21. Classification of plants: vascular and non-vascular
22. Vascular tissues - xylem and phloem
23. Seed, seed coat, cotyledon, seed dispersal, germination
24. Structure and function of leaf, stem and root
25. Angiosperms and gymnosperms
26. Structure and function of flowers - flower parts and the function of each part
27. Pollination
28. Tropism



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Mark Steyvers



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web oxidation reduction redox reactions oxidation $\text{Cu} + \text{S} + \text{Cu} + 2\text{A} + \text{Q} + 2\text{E}$ reduction $\text{MnO}_4 + 4\text{A} + \text{A} + \text{Q} + 4\text{H} + \text{A} + \text{Q} + 3\text{E} + \text{MnO}_4 + 2\text{S} + 2\text{H} + \text{A} + 2\text{O}$ l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} + \text{S} + \text{MnO}_4 + 4\text{A} + \text{A} + \text{Q}$ is

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web 1 practice problem balance the redox reaction in basic solution $\text{NO}_2 + \text{AQ} + \text{Al} + \text{S} + \text{NH}_3 + \text{G} + \text{ALO}_2 + \text{AQ}_2$ practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of OH^- br 2 l bro 3 aq br aq previous topic next topic

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web google classroom for the following redox equation $\text{MnO}_4 + 4\text{A} + \text{AQ} + \text{SO}_2 + 2\text{G} + \text{Mn} + 2\text{AQ} + \text{HSO}_4 + 4\text{A} + \text{AQ}$ what will be the stoichiometric coefficient of HSO_4^- if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web bonus problem here is a balanced net ionic equation $3\text{Br}_2 + 6\text{OH}^- + 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$ the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^{-}$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^{-} + 14\text{H}^{+} + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^{-}$ equalize the electrons $6\text{e}^{-} + 14\text{H}^{+} + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^{-}$ multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add H^{+} ions or the H_2O pair to fully balance the equation in this video we ll walk through this process for the reaction between dichromate $\text{Cr}_2\text{O}_7^{2-}$ and chloride Cl^{-} ions in acidic solution created by jay

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_2 + \text{SO}_4^{2-} + \text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^{-} + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method 8 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^{+} + \text{TeO}_3^{2-} + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te} + 10\text{H}_2\text{O}$ 11 $\text{PbO}_2 + \text{IO}_3^{-} + \text{I}_2 + \text{Pb} + 2\text{H}_2\text{O}$ 12 $\text{AsO}_3^{-} + \text{IO}_3^{-} + \text{H}_2\text{AsO}_4^{-}$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{H}_2\text{NO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ c H_2NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ d $\text{C}_6\text{H}_8\text{O}_6$ and H_2NO_2