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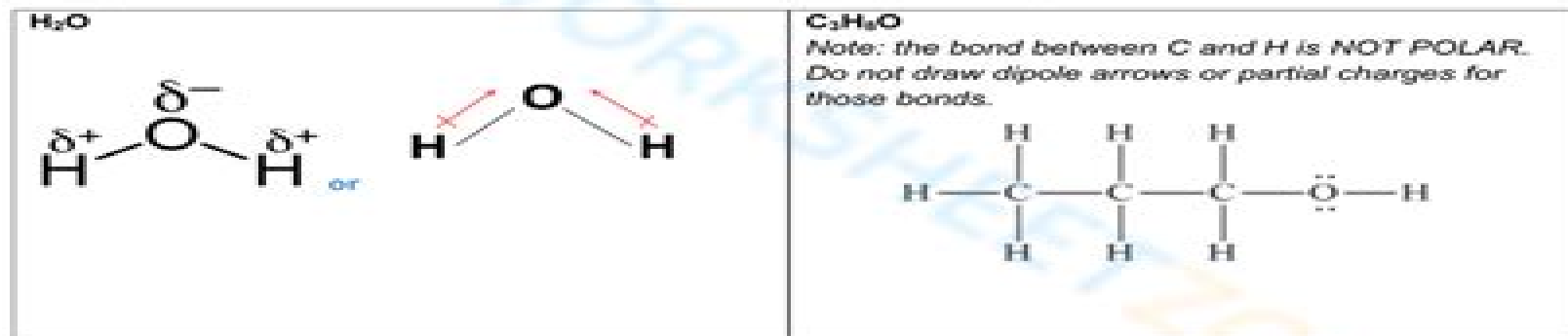
Intermolecular Forces Lab Worksheet - Answer Key



Directions: Today you will complete two experiments that compare the properties of water to isopropanol (a type of rubbing alcohol). Before conducting each experiment, you will first read about the experiment and then make a prediction about what you think will happen. Afterward you do the experiment, you will write an explanation of what was happening at the atomic level that let us observe these properties.

Introduction: In many ways, water is a miracle liquid. It is essential for all living things on earth and it is often referred to as a universal solvent because many substances dissolve in it. Water displays unusual properties due to the ways the individual water molecules interact with each other.

1. Draw the structural formula (with lone pairs) for the two molecules we will be comparing today, water (H_2O) and isopropanol ($\text{C}_3\text{H}_8\text{O}$). Draw a dipole arrow to show the direction the electrons are being pulled and label the areas of partial positive and negative charges.



2. Which molecule is more polar? water

Experiment #1: Penny

How many drops of water fit on the head of a penny? How many drops of isopropanol fit on the head of a penny?

Prediction:

- a. I think _____ drops of water will fit on the head of the penny because _____

Intermolecular Forces Lab And Answers

Peipei Pang



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subject alive many real life examples and applications have been included This book brings together the often scattered and diffuse background information that is essential for a full understanding of chemical modeling FEATURES Assumes no prior knowledge of modeling Brings together all the background information that a reader needs to know to fully understand modeling Develops classical and quantum mechanical theories from basics Avoids unnecessary mathematical rigour Includes a detailed mathematical Toolbox as a ready reference Includes brief descriptions and web addresses of key software packages CONTENTS Introduction Acknowledgements Software Packages Describing Macroscopic Systems Thermodynamics R sum of Classical Mechanics Modeling Simple Solids I Introduction to Quantum Mechanics Electric Multipoles Polarizabilities and Intermolecular Forces Some Statistical Ideas Applications of the Boltzmann Distribution Modeling Simple Solids II Molecular Mechanics Molecular Dynamics and Monte Carlo Techniques The Ideal Monatomic Gas Quantum Gases Introduction to Statistical Thermodynamics Modeling Atoms Diatomics Quantum Modeling of Larger Systems Describing Electron Correlation The Band Theory of Solids Modeling Polymeric Materials Modeling Liquids Appendices Suggestions for Further Reading Index *Aeronautical Engineering Review* ,1955 **Rubber Journal** ,1945 *Physics Briefs* ,1991

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