

KEY

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of K_a for the acid HBrO. (4 marks) $\text{pH} = 4.65$ so $[\text{H}_3\text{O}^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} \text{ M}$

$$\text{HBrO} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{BrO}^-$$

[I]	0.25		0	0
[C]	-2.239×10^{-5}		$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$0.25 - 2.239 \times 10^{-5}$		2.239×10^{-5}	2.239×10^{-5}

④

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{BrO}^-]}{[\text{HBrO}]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks) $\text{pH} = 2.355$ $[\text{H}_3\text{O}^+] = \text{antilog}(-2.355) = 0.004416 \text{ M}$

$$\text{C}_6\text{H}_5\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{C}_6\text{H}_5\text{COO}^-$$

[I]	C_0		0	0
[C]	-0.004416		$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$		0.004416	0.004416

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{C}_6\text{H}_5\text{COO}^-]}{[\text{C}_6\text{H}_5\text{COOH}]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 = 0.30439 \text{ M}$$

Answer $C_0 = 0.30 \text{ M}$

5. Find the value of K_b for the oxalate ion ($\text{C}_2\text{O}_4^{2-}$). (1 mark)

$$K_b(\text{C}_2\text{O}_4^{2-}) = \frac{K_w}{K_a(\text{HC}_2\text{O}_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

①

Answer $K_b = 1.6 \times 10^{-10}$

6. The value of K_b for the weak base methylamine (CH_3NH_2) is 4.4×10^{-4} . Calculate the value of K_a for the acid CH_3NH_3^+ . (1 mark)

$$K_a(\text{CH}_3\text{NH}_3^+) = \frac{K_w}{K_b(\text{CH}_3\text{NH}_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

①

Answer $K_a = 2.3 \times 10^{-11}$

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SA Dillow



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