

# CHEMISTRY

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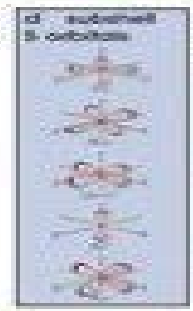
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## PERIODIC TABLE OF THE ELEMENTS

The periodic table is color-coded by groups:
 

- Group 1: s-block, 1 orbital
- Group 2: s-block, 1 orbital
- Groups 13-18: p-block, 5 orbitals
- Groups 3-10: d-block, 5 orbitals
- Groups 11-12: s-block, 1 orbital

 A diagram for Hydrogen (H) shows its atomic structure with 1 proton and 1 electron, labeled as 'atomic number' and 'atomic weight'.



### ATOMIC STRUCTURE

Atomic Number, Z, is of protons in the nucleus. For neutral atom, Z = # of electrons. Atomic weight, A, is sum of # of protons and neutrons in the nucleus. Mass Number, M, is Z + N, where N is number of neutrons in the nucleus. Isotopes of an element with the same Z, different N, are shown by the number of protons and neutrons. A number of isotopes is practical with the element. Mass Number or Atomic Weight is practical average of isotopes for a given element.

### ATOMIC QUANTUM NUMBERS AND ORBITALS

Quantum numbers describe the energy, size, and shape of an atomic orbital. The principal quantum number, n, is the most important. It is the number of shells. Each value of n defines a shell. In the ground state, each shell contains a certain number of orbitals.

n=1	1 orbital	1 shell
n=2	4 orbitals	2 shells
n=3	9 orbitals	3 shells

### NUCLEAR CHEMISTRY

Nuclear reactions occur when nuclear particles (neutrons, protons, alpha particles, beta particles, gamma rays) interact with nuclei. Energy is released or absorbed. Nuclear reactions are balanced for mass and charge. Radioactive decay is the process by which an unstable nucleus loses energy by emitting ionizing radiation. Types of decay: alpha, beta, gamma, positron, electron capture. Half-life is the time required for half of the nuclei in a sample to decay.

### MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in many-electron atoms. The Pauli Exclusion Principle states that no two electrons in an atom can have the same four quantum numbers. The Aufbau Principle states that electrons fill orbitals from lowest to highest energy.

n	l	m <sub>l</sub>	m <sub>s</sub>	Orbital
1	0	0	±1/2	1s orbital
2	0	0	±1/2	2s orbital
2	1	0, ±1	±1/2	2p orbitals
3	0	0	±1/2	3s orbital
3	1	0, ±1	±1/2	3p orbitals
3	2	0, ±1, ±2	±1/2	3d orbitals

Electron configuration: 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>4</sup>

A detailed periodic table showing element symbols, atomic numbers, and group numbers. The table is color-coded to match the atomic structure diagram.

# Chemistry Quickstudy Reference Guides Academic

**Shasha Hu**



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