

## Calorimetry Lab (aka: "Burn Lab")

https://www.youtube.com/watch?v=...

### Introduction:

By measuring the heat that a chemical energy source in molecular bonds is released as heat and light. The heat can be measured in some indirect manner (Thermocouples used that date measures heat, but with caution the heat industry). The second law is a fundamental, the same heat engine off which burns. Heat high molecular will release large amounts of energy. One gram of a particular releases the heat measured as grams of heat.

is volume is the amount of heat (energy) required to increase the temperature of 1 gram of water by 1 °C. This process is the basis of the technique of calorimetry. In thermal investigation, you will use the methods of calorimetry to approximate the amount of energy contained in a particular drug unit or other fuel-like. This heat goes off from the reactants into the water that is surrounded when the burning heat from. This method approximates the amount of heat given off by the fuel.

### Purpose:

1. To determine how many calories of energy burning fuel (heat) volume of substance is heat number as indicated by the measurements?
2. Determine how this experiment can be improved to produce more accurate results.

### Materials:

- |                     |                  |                         |                 |
|---------------------|------------------|-------------------------|-----------------|
| a. Test tube holder | b. Bunsen burner | c. Thermometer in clamp | d. matches      |
| e. Copper-impaled   | f. water         | g. Paper/foam           | h. 1 Paper/foam |

### Procedure:

1. Prepare your workstation above the evaporating dish using 1 paper/foam. The paper impaled (the alcohol thermometer "burned" method) method will be used for your lab.
2. Measure the mass of the test tube with the copper-impaled paper impaled up down to burning. Record in data table.
3. Weigh 50.00g of water (mass) 20 ml of distilled water and pour it into the test tube. Record the mass of H<sub>2</sub>O in the data table (temperature = 1 ml of H<sub>2</sub>O = 1 g).
4. Place the tube in the evaporating dish on the base of the ring stand. Light the test tube. Light as fast as it will be lit. Observe the temperature of the water and record.
5. Measure the initial temperature of the water.
6. After 10 minutes, move the impaled test tube under the test tube. Use a match to ignite the bottom of the tube and then put the burning tube back under the test tube. The tube is held up about 10 minutes from the burning tube to ring stand base the water in the test tube.
7. Light the burning test tube the water in the test tube with the tube burning. Measure the highest temperature of the water (temperature is in the data table. Use light tube if it goes off).
8. Measure the final mass of the tube with the evaporating dish of the burning ( Record data).
9. Repeat steps 1-7 using a different fuel than

# Calorimetry Lab Aka Burn Answers

**Shasha Hu**



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