

Half Modern Chemistry Review

CHAPTER 9: STOICHIOMETRY

The following pages contain the best (but not all) of the information for the chapter 9 test.

Focus on this content, but make sure to review class notes, activities, worksheets, questions, etc.

If you study this document and NOTHING else, you should at least be able to PASS the test.

**** Test items will be recall, recognition, and/or application of this content. ****

OBJECTIVES

- Calculate molar weight for molecular-formula compounds. (C 1)
- Calculate molar weight for ionic compounds. (C 1)
- 9.1. Determine moles of reactants from balanced chemical equations. (T & P)
- 9.2. Perform stoichiometry calculations with: mol to mol, mol to gram, gram to mol, and gram to gram. (P & P)
- 9.3. Calculate stoichiometry problems and determine the limiting reagent, excess reagent, and amount of product produced. (P & P)

9.1: INTRODUCTION TO STOICHIOMETRY

a. Introductory

- composition stoichiometry – calculations involving the mass relationships of elements in compounds.
- reaction stoichiometry – calculations involving the mass relationships between reactants and products in a chemical reaction.
- molar ratio – a conversion factor that relates the amount, in moles, of any two substances involved in a chemical reaction.

b. Chapter Highlights

- Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.
- Stoichiometry is based on the law of conservation of mass.
- Reacting mass substances is specified using the amount of each substance in moles.
- A molar ratio in the chemical formula that relates the amount in moles of any two substances in a chemical reaction. This molar ratio is derived from the balanced equation.
 - Example: $2\text{FeCl}_3 + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_4$ which is the ratio of iron to chlorine moles? $2/3$ or $3/2$
- Amount of a substance is measured in moles, and mass of a substance is represented using mass units such as grams, kilograms, or milligrams.
- Mass and amount of substances are quantities, whereas moles and grams are units.
- A balanced chemical equation is necessary to solve any stoichiometric problem.

9.2: BASIC STOICHIOMETRIC CALCULATIONS

a. Chapter Highlights

- In an ideal stoichiometric calculation, the mass of the amount of any reactant or product can be calculated if the balanced chemical equation and the molar mass (amount of any unit) is known, as outlined in lesson.

9.3: LIMITED REACTANTS AND PERCENTAGE YIELD

Modern Chemistry Chapter 9 Section 1 Review Answers

CL Gary



Modern Chemistry Chapter 9 Section 1 Review Answers:

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