

**KEY**

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of Ka for the acid HBrO. (4 marks)  $pH = 4.65$  so  $[H_3O^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} M$



[I]	0.25		0	0
[C]	$-2.239 \times 10^{-5}$		$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$0.25 - 2.239 \times 10^{-5}$		$2.239 \times 10^{-5}$	$2.239 \times 10^{-5}$

4

$$K_a = \frac{[H_3O^+][BrO^-]}{[HBrO]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer  $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks)  $pH = 2.355$   $[H_3O^+] = \text{antilog}(-2.355) = 0.004416 M$



[I]	$C_0$		0	0
[C]	$-0.004416$		$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$		$0.004416$	$0.004416$

4

$$K_a = \frac{[H_3O^+][C_6H_5COO^-]}{[C_6H_5COOH]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 = 0.30439 M$$

Answer  $C_0 = 0.30 M$

5. Find the value of Kb for the oxalate ion ( $C_2O_4^{2-}$ ). (1 mark)

$$K_b(C_2O_4^{2-}) = \frac{K_w}{K_a(HC_2O_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

1

Answer  $K_b = 1.6 \times 10^{-10}$

6. The value of Kb for the weak base methylamine ( $CH_3NH_2$ ) is  $4.4 \times 10^{-4}$ . Calculate the value of Ka for the acid  $CH_3NH_3^+$ . (1 mark)

$$K_a(CH_3NH_3^+) = \frac{K_w}{K_b(CH_3NH_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

1

Answer  $K_a = 2.3 \times 10^{-11}$

10

# Basic Stoichiometry Phet Lab Answer Key

**S Marginson**



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web figure 27 2 male reproductive system the structures of the male reproductive system include the testes the epididymides the penis and the ducts and glands that produce and carry semen sperm exit the scrotum through the ductus deferens which is bundled in the spermatic cord

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