

**KEY**

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of Ka for the acid HBrO. (4 marks)  $pH = 4.65$  so  $[H_3O^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} M$



[I]	0.25		0	0
[C]	$-2.239 \times 10^{-5}$		$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$0.25 - 2.239 \times 10^{-5}$		$2.239 \times 10^{-5}$	$2.239 \times 10^{-5}$

4

$$K_a = \frac{[H_3O^+][BrO^-]}{[HBrO]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer  $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks)  $pH = 2.355$   $[H_3O^+] = \text{antilog}(-2.355) = 0.004416 M$



[I]	$C_0$		0	0
[C]	$-0.004416$		$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$		$0.004416$	$0.004416$

4

$$K_a = \frac{[H_3O^+][C_6H_5COO^-]}{[C_6H_5COOH]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 = 0.30439 M$$

Answer  $C_0 = 0.30 M$

5. Find the value of Kb for the oxalate ion ( $C_2O_4^{2-}$ ). (1 mark)

$$K_b(C_2O_4^{2-}) = \frac{K_w}{K_a(HC_2O_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

1

Answer  $K_b = 1.6 \times 10^{-10}$

6. The value of Kb for the weak base methylamine ( $CH_3NH_2$ ) is  $4.4 \times 10^{-4}$ . Calculate the value of Ka for the acid  $CH_3NH_3^+$ . (1 mark)

$$K_a(CH_3NH_3^+) = \frac{K_w}{K_b(CH_3NH_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

1

Answer  $K_a = 2.3 \times 10^{-11}$

10

# Basic Stoichiometry Phet Lab Answer Key

**B Lingard**



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### **Table of Contents Basic Stoichiometry Phet Lab Answer Key**

1. Understanding the eBook Basic Stoichiometry Phet Lab Answer Key
  - The Rise of Digital Reading Basic Stoichiometry Phet Lab Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Basic Stoichiometry Phet Lab Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Basic Stoichiometry Phet Lab Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Stoichiometry Phet Lab Answer Key
  - Personalized Recommendations

- Basic Stoichiometry Phet Lab Answer Key User Reviews and Ratings
- Basic Stoichiometry Phet Lab Answer Key and Bestseller Lists
- 5. Accessing Basic Stoichiometry Phet Lab Answer Key Free and Paid eBooks
  - Basic Stoichiometry Phet Lab Answer Key Public Domain eBooks
  - Basic Stoichiometry Phet Lab Answer Key eBook Subscription Services
  - Basic Stoichiometry Phet Lab Answer Key Budget-Friendly Options
- 6. Navigating Basic Stoichiometry Phet Lab Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Basic Stoichiometry Phet Lab Answer Key Compatibility with Devices
  - Basic Stoichiometry Phet Lab Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Basic Stoichiometry Phet Lab Answer Key
  - Highlighting and Note-Taking Basic Stoichiometry Phet Lab Answer Key
  - Interactive Elements Basic Stoichiometry Phet Lab Answer Key
- 8. Staying Engaged with Basic Stoichiometry Phet Lab Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Basic Stoichiometry Phet Lab Answer Key
- 9. Balancing eBooks and Physical Books Basic Stoichiometry Phet Lab Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Basic Stoichiometry Phet Lab Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Basic Stoichiometry Phet Lab Answer Key
  - Setting Reading Goals Basic Stoichiometry Phet Lab Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Basic Stoichiometry Phet Lab Answer Key
  - Fact-Checking eBook Content of Basic Stoichiometry Phet Lab Answer Key

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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