

Chapter 17: Chemical Equilibrium

Worksheet: K<sub>c</sub> and K<sub>p</sub>

1. For the reaction  $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$

at 1000 K

- 1. The equilibrium constant  $K_c$  for the reaction  $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$  is 16.0 at 1000 K. Calculate  $K_c$  for the reaction  $\text{SO}_2(g) + \frac{1}{2}\text{O}_2(g) \rightleftharpoons \text{SO}_3(g)$  at 1000 K.
- 2. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at 400 K. Calculate  $K_c$  for the reaction  $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$  at 400 K.
- 3. For the reaction  $\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ ,  $K_c = 4.0$  at 1000 K. Calculate  $K_c$  for the reaction  $\text{NO}(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{1}{2}\text{O}_2(g)$  at 1000 K.
- 4. For the reaction  $\text{CO}(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}_2(g) + \text{H}_2(g)$ ,  $K_c = 1.0$  at 1000 K. Calculate  $K_c$  for the reaction  $2\text{CO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{CO}_2(g)$  at 1000 K.
- 5. For the reaction  $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$ ,  $K_c = 50.0$  at 400 K. Calculate  $K_c$  for the reaction  $\text{HI}(g) \rightleftharpoons \frac{1}{2}\text{H}_2(g) + \frac{1}{2}\text{I}_2(g)$  at 400 K.
- 6. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at 400 K. Calculate  $K_c$  for the reaction  $\text{NH}_3(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g)$  at 400 K.
- 7. For the reaction  $\text{CO}(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}_2(g) + \text{H}_2(g)$ ,  $K_c = 1.0$  at 1000 K. Calculate  $K_c$  for the reaction  $2\text{CO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{CO}_2(g)$  at 1000 K.
- 8. For the reaction  $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2\text{HI}(g)$ ,  $K_c = 50.0$  at 400 K. Calculate  $K_c$  for the reaction  $\text{HI}(g) \rightleftharpoons \frac{1}{2}\text{H}_2(g) + \frac{1}{2}\text{I}_2(g)$  at 400 K.
- 9. For the reaction  $\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ ,  $K_c = 0.50$  at 400 K. Calculate  $K_c$  for the reaction  $\text{NH}_3(g) \rightleftharpoons \frac{1}{2}\text{N}_2(g) + \frac{3}{2}\text{H}_2(g)$  at 400 K.
- 10. For the reaction  $\text{CO}(g) + \text{H}_2\text{O}(g) \rightleftharpoons \text{CO}_2(g) + \text{H}_2(g)$ ,  $K_c = 1.0$  at 1000 K. Calculate  $K_c$  for the reaction  $2\text{CO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{CO}_2(g)$  at 1000 K.

Answers

- 1.  $K_c = 4.0$
- 2.  $K_c = 2.0$
- 3.  $K_c = 0.25$
- 4.  $K_c = 1.0$
- 5.  $K_c = 0.020$
- 6.  $K_c = 0.25$
- 7.  $K_c = 1.0$
- 8.  $K_c = 0.020$
- 9.  $K_c = 0.25$
- 10.  $K_c = 1.0$

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