

CHEMISTRY PRACTICAL - SALT ANALYSIS

NOTE:

1. Exp. 1 and 2 for your reference, no need to write in your notebook.
2. Exp. 3 to 10 are to be written on the given table.
3. Chemical Equations are to be learned by the students thoroughly.
4. Start each Experiment from new page.

EXPERIMENT – 1 Classification of Anions

Group	Group Reagent	Observation	Inference
A	Dilute H_2SO_4	a) Colourless, odourless gas with brisk effervescence (CO_2) which turns lime water milky. b) Colourless gas with rotten egg like smell (H_2S) which turns lead acetate paper black. c) Colourless gas with smell of burning sulphur (SO_2) which turns acidified dichromate paper green. d) Brown coloured gas (NO_2) which turns ferrous sulphate solution black or brown. e) Colourless gas with vinegar like smell.	CO_3^{2-} (Carbonate) S^{2-} (sulphide) SO_3^{2-} (Sulphite) NO_2^- (Nitrite) CH_3COO^- (Acetate)
B	Conc. H_2SO_4	a) Colourless pungent smelling gas (HCl) which gives white dense fumes with glass rod dipped in NH_4OH . b) Violet coloured vapours (I_2) which turns starch paper blue. c) Reddish brown gas (NO_2) having pungent smell (On adding copper turning, fumes becomes intense) d) Brown colour gas with pungent smell (Br_2) which turns starch paper yellow. e) Colourless, odourless gas with brisk effervescence ($CO + CO_2$) which turns lime water milky and burns on the mouth of test tube with blue flame.	Cl^- (Chloride) I^- (Iodide) NO_3^- (Nitrate) Br^- (Bromide) $C_2O_4^{2-}$ (Oxalate)
C	$BaCl_2$	White ppt. of $BaSO_4$ is formed.	SO_4^{2-} (sulphate)
D	Ammonium molybdate $3(NH_4)_2MoO_4$	Cannary yellow ppt. of phospho ammonium molybdate ($(NH_4)_3PO_4 \cdot 12 MoO_5 \cdot 6H_2O$)	PO_4^{3-} (phosphate)

EXPERIMENT – 2

Classification of Cations

Group	Group Reagent	Radical	PPt/Smell	Colour
Zero	$NaOH$	NH_4^+	Smell of NH_3	-
I	Dil. HCl	Pb^{2+}	$PbCl_2$	White
II	H_2S gas in acidic Medium	Pb^{2+} As^{3+} Cu^{2+} Cd^{2+}	PbS As_2S_3 CuS CdS	Black Yellow Black Yellow
III	NH_4Cl (s) in presence of NH_4OH	Fe^{3+} Fe^{2+} Al^{3+}	$Fe(OH)_3$ $Fe(OH)_2$ $Al(OH)_3$	Light green Reddish brown Gelatinous white
IV	H_2S gas in basic medium	Ni^{2+}	NiS	Black

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