

Answers to Chapter 3 Study Questions

1. a) $3(12.0) + 8(1.01) + 3(16.0) = 92.1 \text{ g/mole}$ b) 92.1 g c) 6.02×10^{23} molecules
 d) $0.217 \text{ moles} \times \frac{92.1 \text{ g}}{1 \text{ mole}} = 20.0 \text{ g}$
2. a) 4 atoms (one N + 3 H)
 b) $1 \text{ mole } \text{NH}_3 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 2.41 \times 10^{24} \text{ atoms}$
 c) $3.40 \text{ g } \text{NH}_3 \times \frac{1 \text{ mole } \text{NH}_3}{17.0 \text{ g } \text{NH}_3} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole } \text{NH}_3} \times \frac{4 \text{ atoms}}{1 \text{ molecule}} = 4.82 \times 10^{23} \text{ atoms}$
3. Molar mass of $\text{NaNO}_2 = 23.0 + 14.0 + 2(16.0) = 69.0 \text{ g/mole}$
 $\% \text{ Na} = 23.0/69.0 = 33.3\% \text{ Na}$; $\% \text{ N} = 14.0/69.0 = 20.3\% \text{ N}$; $\% \text{ O} = 2(16.0)/69.0 = 46.4\% \text{ O}$
33.3% Na, 20.3% N and 46.4% O.
4. a) In 100 g of this compound, there are 40.7 g C, 5.1 g H, and 54.2 g O
 $40.7 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 3.39 \text{ moles C}$ $3.39/3.39 = 1 \times 2 = 2$
 $5.1 \text{ g H} \times \frac{1 \text{ mole H}}{1.0 \text{ g H}} = 5.1 \text{ moles H}$ $5.1/3.39 = 1.5 \times 2 = 3$
 $54.2 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 3.39 \text{ moles O}$ $3.39/3.39 = 1 \times 2 = 2$
 empirical formula = **$\text{C}_2\text{H}_3\text{O}_2$**
- b) Molar mass of $\text{C}_2\text{H}_3\text{O}_2 = 2(12.0) + 3(1.0) + 2(16.0) = 59.0 \text{ g/mole}$
 $118/59.0 = 2 \rightarrow$ molecular formula = **$\text{C}_4\text{H}_6\text{O}_4$**
5. a) In 25.0 g of compound, there are 7.20 g Mg, 3.55 g C and $25.0 - (7.20 + 3.55) = 14.25 \text{ g O}$.
 $7.20 \text{ g Mg} \times \frac{1 \text{ mole Mg}}{24.3 \text{ g Mg}} = 0.296 \text{ moles Mg}$ $0.296/0.296 = 1$
 $3.55 \text{ g C} \times \frac{1 \text{ mole C}}{12.0 \text{ g C}} = 0.296 \text{ moles C}$ $0.296/0.296 = 1$
 $14.25 \text{ g O} \times \frac{1 \text{ mole O}}{16.0 \text{ g O}} = 0.891 \text{ moles O}$ $0.891/0.296 = 3$
 formula = **MgCO_3**
- b) $\% \text{ Mg} = 7.20/25.0 = 28.8\% \text{ Mg}$; $\% \text{ C} = 3.55/25.0 = 14.2\% \text{ C}$; $\% \text{ O} = 14.25/25.0 = 57.0\% \text{ O}$. (You should get the same result using molar mass and atomic masses.)
- c) $13.9 \text{ g compound} \times \frac{28.8 \text{ g Mg}}{100 \text{ g compound}} = 4.00 \text{ g Mg}$

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ML Morrison



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