

Science

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(Chapter – 1) (Chemical Reactions and Equations) (Notes)

(Class – X)

Thermal Decomposition:

When a decomposition reaction is carried out by heating, it is called thermal decomposition.

Decomposition of Silver chloride:



white silver chloride turns grey in sunlight. This is due to the decomposition of silver chloride into silver and chlorine by light.

- This Reaction is used in black and white photography.
- All decomposition reactions require energy either in the form of heat, light or electricity for breaking down reactants.

On the basis of heat there are two kinds of reaction:

A). Exothermic Reactions:

Reactions in which heat is released along with the formation of products are called exothermic.

During the reaction large amount of heat (energy) is evolved

This heat makes the reaction mixture warm.

Examples of exothermic reactions:

(i) Burning of natural gas



(ii) Respiration is also an exothermic reaction.

When we take food to get energy to stay alive. During digestion, food is broken down into simpler substances like carbohydrates and other nutrients. These carbohydrates further broken down to form glucose. This glucose combines with oxygen in cells of our body in the process of cellular respiration and provide energy. Therefore, respiration is also an exothermic reaction.

(iii) The decomposition of vegetable matter into compost is also an example of an exothermic reaction.

B). Endothermic Reactions: Reactions in which energy is absorbed are known as endothermic reactions.

3. Displacement Reaction

The reaction in which the highly reactive element displaces the less reactive element from the reaction, such reaction is known as **displacement reaction**.

General formula for displacement reaction:



Here C is highly reactive element which displace A form the reaction and joins b.

Example:

When a iron nail is placed in the aqueous solution of copper sulphate Iron displaces copper from the reaction or its compound and blue colour of copper sulphate turns grey brown because Iron is highly reactive metal than copper.

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1. Understanding the eBook Class 10 Science Notes Nepal Download
 - The Rise of Digital Reading Class 10 Science Notes Nepal Download
 - Advantages of eBooks Over Traditional Books
2. Identifying Class 10 Science Notes Nepal Download
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Class 10 Science Notes Nepal Download
 - User-Friendly Interface
4. Exploring eBook Recommendations from Class 10 Science Notes Nepal Download
 - Personalized Recommendations
 - Class 10 Science Notes Nepal Download User Reviews and Ratings
 - Class 10 Science Notes Nepal Download and Bestseller Lists

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 - Highlighting and Note-Taking Class 10 Science Notes Nepal Download
 - Interactive Elements Class 10 Science Notes Nepal Download
8. Staying Engaged with Class 10 Science Notes Nepal Download
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Class 10 Science Notes Nepal Download
9. Balancing eBooks and Physical Books Class 10 Science Notes Nepal Download
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Class 10 Science Notes Nepal Download
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Class 10 Science Notes Nepal Download
 - Setting Reading Goals Class 10 Science Notes Nepal Download
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Class 10 Science Notes Nepal Download
 - Fact-Checking eBook Content of Class 10 Science Notes Nepal Download
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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