

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, which are exploring the structures through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
4. In part 1, what is the purpose of adding potassium thiocyanate to 100 mL of 0.1 M Fe^{3+} solution? What is the color of the solution?
The purpose of adding thiocyanate to the solution is to form a complex with iron(III) ions. The color of the solution is a deep red color.
5. What is the purpose of adding distilled water to part 1?
Water is added to dilute the solution, which causes the color to become lighter.
6. What is the purpose of adding 6 M HCl to part 1?
6 M HCl is added to increase the concentration of H^{+} ions, which causes the color to become darker.
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$
8. Write your observations for [part 7].
Adding 6 M OH^{-} to the solution causes the formation of a precipitate, which is a dark brown color.
9. What is the purpose of the reaction in part 7?
The purpose of the reaction is to form a precipitate, which is a dark brown color.

Investigating Chemical Equilibrium Lab 12a Answer

Katrin Zwirgmaier



Investigating Chemical Equilibrium Lab 12a Answer:

The Enigmatic Realm of **Investigating Chemical Equilibrium Lab 12a Answer**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Investigating Chemical Equilibrium Lab 12a Answer** a literary masterpiece penned by a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting impact on the hearts and minds of people who partake in its reading experience.

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