

Pre-lab: Acid-Base Titrations

CHM205 Laboratory

Fall 2019 Section: _____

Grade: _____

Name: _____

A procedure similar to the one in this lab is used to titrate a sodium hydroxide solution against potassium hydrogen phthalate (KHP). The follow data is recorded: For the standard KHP solution, 10.575 g of KHP is dissolved in water, and diluted to 250.00 mL in a volumetric flask; 20.00 mL of the KHP solution are added to an Erlenmeyer flask; a buret is filled with a NaOH solution of unknown concentration to a starting volume of 0.18 mL, and at the end point of the titration the buret reading is 20.83 mL.

1. Calculate the mmoles KHP that were put into the 250.00 mL volumetric flask. Note: KHP's chemical formula is $\text{KHC}_8\text{H}_4\text{O}_4$.
2. Calculate the KHP solution concentration (in molarity units) that was prepared in the 250 mL volumetric flask. Note sig figs on the flask volume.
3. Calculate the mmoles KHP that are in the 20.00 mL aliquot.
4. Write the balanced chemical equation for the reaction of KHP with sodium hydroxide.
5. Calculate the mmoles NaOH that will be needed to reach the titration end point, when the 20.00 mL aliquot is completely titrated.
6. Calculate the volume of NaOH used in the titration.
7. Calculate the concentration (in molarity units) of the NaOH solution.

Acid Base Titrations Pre Lab Answers

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Acid Base Titrations Pre Lab Answers:

Integrated Approach to Coordination Chemistry Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07

Coordination chemistry is the study of compounds formed between metal ions and other neutral or negatively charged molecules. This book offers a series of investigative inorganic laboratories approached through systematic coordination chemistry. It not only highlights the key fundamental components of the coordination chemistry field, it also exemplifies the historical development of concepts in the field. In order to graduate as a chemistry major that fills the requirements of the American Chemical Society, a student needs to take a laboratory course in inorganic chemistry. Most professors who teach and inorganic chemistry laboratory prefer to emphasize coordination chemistry rather than attempting to cover all aspects of inorganic chemistry because it keeps the students focused on a cohesive part of inorganic chemistry which has applications in medicine, the environment, molecular biology, organic synthesis, and inorganic materials.

Chemindustry Experiments Brenda W. Hill, 1979 *Project SERAPHIM 1991 Catalog*, 1991 **Using Physical Models of Biomolecules to Teach**

Concepts of Biochemical Structure in Introductory Undergraduate Chemistry John Yi, 2004 *Chemical*

Investigations Nancy Konigsberg Kerner, 1986 **Chemistry in the Laboratory** James M. Postma, Julian L. Roberts, J.

Leland Hollenberg, 2009-08-07 For nearly 40 years Chemistry in the Laboratory has been meeting the needs of teachers and students. This new edition builds on that legacy while addressing cutting edge trends in the chemistry laboratory including forensic chemistry and environmental and green chemistry. As always, the new edition of Chemistry in the Laboratory offers precise, easy-to-follow instructions, helpful illustrations, and an emphasis throughout on laboratory safety. Again throughout, a Consider This feature encourages users to expand the principles of the experiment into interesting applications, open-ended experiments, or unexplored corners. Most experiments in the manual can be completed in one lab session, but some can be linked or extended for a multi-lab project.

Biology/science Materials Carolina Biological Supply Company, 1991

Selected Water Resources Abstracts, 1974-03 **Instructor's Manual** Brian F. Woodfield, Matthew C. Asplund, 2006

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Journals in Education, 1972 **Analysis Instrumentation**, 1980 Contains Proceedings of the annual Analysis

Instrumentation Symposium Teaching Geology Using the History and Philosophy of Science Glenn Dolphin, 2024-10-02

This book provides a case study on how to design and build an introductory geology course for non science majors. The book presents a foundation with the status of geoscience education and research in geoscience conceptual development as a backdrop for the design process. It then describes the instructional goal setting process and development of the structural components of the course based on the determined goals. The book presents the three historical narratives: the earth is a historical entity, the earth is very old, and the earth is dynamic, that form the foundation of instruction. It also describes examples of the implicit, explicit, and reflective treatments of the nature of science to help students develop a better sense of the process of geology. Finally, the book gives preliminary results from some innovative approaches to research on student learning within the domains of geological content knowledge and NOS content knowledge within the course.

Energy Research Abstracts, 1980 Current Chemical Papers Chemical Society (Great Britain), 1964 A classified world list of new papers in pure chemistry *Chemistry in Canada*, 1974 **The Software Encyclopedia**, 1988 Electroanalytical Abstracts, 1978 International journal dealing with the documentation of all aspects of fundamental physico-chemical and analytical electrochemistry **Laboratory Practice**, 1988

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