

## Alberta Exam Chem 30 unit 4 Version 2 ANSWER KEY

1. In the acid-base reaction



- A) reactants are favoured  
 B)  $\text{HC}_6\text{H}_5\text{O}_2^-(\text{aq})$  is acting as an acid  
 C)  $\text{SO}_4^{2-}(\text{aq})$  is a proton acceptor  
 D)  $\text{HSO}_4^-(\text{aq})$  and  $\text{HC}_6\text{H}_5\text{O}_2^-(\text{aq})$  are a conjugate acid-base pair

**Incorrect.** Your answer=A, Correct answer=C

**Explanation:** In the above reaction, the conjugate acid base pair is  $\text{HSO}_4^-$  (acting as an acid, proton donor) and  $\text{SO}_4^{2-}$  (conjugate base, proton acceptor). Choice C is a correct answer.

- 2.

Tooth decay results from the dissolving of tooth enamel, or hydroxyapatite,  $\text{Ca}_5(\text{PO}_4)_3\text{OH}(\text{s})$ . Bacteria in the presence of sugars in the mouth change the pH of the mouth which causes the following equilibrium:



When sugar ferments on teeth and gums, the mouth becomes more acidic. Tooth enamel

- A) is formed as the concentration of the hydroxide ion decreases and the equilibrium shifts to the left  
 B) is formed as the concentration of the hydroxide ion increases and the equilibrium shifts to the left  
 C) breaks down as the concentration of the hydroxide ion decreases and the equilibrium shifts to the right  
 D) breaks down as the concentration of the hydroxide ion increases and the equilibrium shifts to the right

**Correct.** Your answer=C, Correct answer=C

**Explanation:** Using your common sense and remembering your last visit to the dentist can help you out here.  $\text{Ca}_5(\text{PO}_4)_3\text{OH}(\text{s})$ , tooth enamel, breaks down in an acid environment. The  $\text{H}^+$  ions would neutralize the  $\text{OH}^-$  ions and the decrease would shift the equilibrium to the right, breaking down the enamel. Only Choice C is correct.

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