NCERT Solution for Class 12
Chemistry
Chapter 3 - Chemical Kinetics

Intext Exercise

1. For the reaction $R \rightarrow P$, the concentration of a reactant changes from 0.03 M to 0.02 M in 25 minutes. Calculate the average rate of reaction using units of time both in minutes and seconds.

Ans: The average rate of the reaction can be calculated by dividing the change in the rate of decreasing the rate of reactant by the time taken. This is given below:

$$\text{Average rate} = -\frac{d[R]}{dt}$$

This can be written as:

$$\text{Average rate} = -\frac{[R]_2 - [R]_1}{t_2 - t_1}$$

$R_1 = 0.03$ M

$R_2 = 0.02$ M

$t_2 - t_1 = 25$ min

Putting the values, we get:

$$\text{Average rate} = -\frac{0.02 - 0.03}{25}$$

$$\text{Average rate} = -\frac{-0.01}{25} = 4 \times 10^{-4}\ \text{M min}^{-1}$$

So, the average rate in minutes will be 4×10^{-4} M min^{-1}

Now, to find the average rate in seconds we have to divide the above answer by 60. So, the answer will be:

$$\text{Average rate} = -\frac{4 \times 10^{-4}}{60} = 6.66 \times 10^{-6}\ \text{M s}^{-1}$$

Therefore, the average rate in seconds will be 6.66×10^{-6} M s^{-1}

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