

Temperature

°C, K

measure of average KE
of motion of particles

Heat

kJ, kcal (Cal) [1 kcal = 4.184 kJ]

measure of total energy transferred from
an object of high E → low E.

Note: A change in T is accompanied by a transfer of heat energy.

Specific heat (c or c_p):

(specific for each substance)

amount of energy required to raise the temperature of 1 g of that substance by 1°C.

change in heat = specific heat x mass x change in T

$$q = c_p \times m \times \Delta T$$

Units: heat absorbed or released, in J;

ΔT in °C or K; m in g; c in J/g°C

Molar heat capacity = heat capacity in J/mol°C instead of J/g°C

Don't forget: the size of 1°C = 1 K, so K can be used in this calculation in place of °C.

Examples

- How much heat energy is released to your body when a cup of hot tea containing 200. g of water is cooled from 65.0°C to body temperature, 37.0°C? (c_p of H₂O = 4.18 J/g°C)
- A 4.50-g nugget of pure gold absorbed 276 J of heat. What was the final temperature of the gold if the initial temperature was 25.0°C? (c_p of gold = 0.129 J/g°C)
- A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance?

Chapter 6 Thermochemistry Test

George M. Bodner, Harry L. Pardue



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