



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{CO}_2$
  - $\text{CaCl}_2$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}_2$
- Write the formula for the following compounds:
  - nitrogen dioxide
  - carbon dioxide
  - hydrogen peroxide
  - carbon monoxide
  - hydrogen cyanide
  - hydrogen sulfide
  - hydrogen chloride
  - hydrogen bromide
  - hydrogen iodide

- Write the molecular formula for each of the following compounds:
  - nitrogen dioxide
  - carbon dioxide
  - hydrogen peroxide
  - carbon monoxide
  - hydrogen cyanide
  - hydrogen sulfide
  - hydrogen chloride
  - hydrogen bromide
  - hydrogen iodide
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  - nitrogen dioxide
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  - hydrogen sulfide
  - hydrogen chloride
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### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride.  $\text{H}-\text{F}$
- Draw the Lewis structure for carbon dioxide.  $\text{O}=\text{C}=\text{O}$
- Draw the Lewis structure for methane.  $\text{C}$  with four single bonds to  $\text{H}$ .
- Draw the Lewis structure for ethane.  $\text{H}-\text{C}-\text{C}-\text{H}$  with single bonds to  $\text{H}$ .
- Draw the Lewis structure for ethyne.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw the Lewis structure for acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$



### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$
- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$

# Chapter 8 Covalent Bonding Practice Problems Answers

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## **Chapter 8 Covalent Bonding Practice Problems Answers:**

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