

CHEM 110 Exam 1 - Practice Test 1 Solutions

1D

The number of protons is always equal to the Atomic number which is the number found on the periodic table above the symbol. The mass number is the number of neutrons added to the number of protons ($\text{Mass} = \#p^+ + \#n^0$). In this problem the mass number is 80. We know that the number of protons is equal to 35, so the number of neutrons must be 45 ($80 = 35p^+ + \#n^0$). The charge on an atom is the number of protons minus the number of electrons. ($\text{Charge} = \#p^+ - \#e^-$). Here, we know the charge is -1 so $(-1 = 35 - \#e^-) = 36$. This makes sense because we would assume that there should be more electrons since the overall charge is negative.

2B

s orbitals are regular circles like A. p orbitals are 8 shaped like B. D orbitals are shaped like answer D. You will learn later in class about hybrid orbitals, which can look like answer C. Bromine on the periodic table is found in the 3-8 group. We know that these groups are filled with electron in the p orbitals. The last electron in the Bromine atom is part of a p orbital so the answer is B.

3C

Keep in mind that for elements other than hydrogen, the 4s is always filled before the 3d. The electron configuration shows the valence electrons to be in the 4 row/family of the periodic table. We see that the s orbital is filled and the d orbital has 8 electrons in it. The only choice is Ni. Answers A and D are tricks. They both have the same number of electrons but different configuration. Remember that when ionizing an atom other than Hydrogen, always take from the highest energy level, not the last filled.

4D

The l quantum number refers to the s orbital. Fe^{3+} has 0 valence electrons in the s orbital, but it has 6 in the core electrons. Keep in mind that electrons in the 4s are ionized before 3d.

5C

This question is based off a periodic trend. Reactivity increases as you go down a group. Therefore this answer is based on simply going down the correct group. Group 1 is the left most group.

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The Princeton Review



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Table of Contents Chem 110 Chapter 1 Practice Test Questions

1. Understanding the eBook Chem 110 Chapter 1 Practice Test Questions
 - The Rise of Digital Reading Chem 110 Chapter 1 Practice Test Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem 110 Chapter 1 Practice Test Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem 110 Chapter 1 Practice Test Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 110 Chapter 1 Practice Test Questions

- Personalized Recommendations
 - Chem 110 Chapter 1 Practice Test Questions User Reviews and Ratings
 - Chem 110 Chapter 1 Practice Test Questions and Bestseller Lists
5. Accessing Chem 110 Chapter 1 Practice Test Questions Free and Paid eBooks
- Chem 110 Chapter 1 Practice Test Questions Public Domain eBooks
 - Chem 110 Chapter 1 Practice Test Questions eBook Subscription Services
 - Chem 110 Chapter 1 Practice Test Questions Budget-Friendly Options
6. Navigating Chem 110 Chapter 1 Practice Test Questions eBook Formats
- ePub, PDF, MOBI, and More
 - Chem 110 Chapter 1 Practice Test Questions Compatibility with Devices
 - Chem 110 Chapter 1 Practice Test Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chem 110 Chapter 1 Practice Test Questions
 - Highlighting and Note-Taking Chem 110 Chapter 1 Practice Test Questions
 - Interactive Elements Chem 110 Chapter 1 Practice Test Questions
8. Staying Engaged with Chem 110 Chapter 1 Practice Test Questions
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem 110 Chapter 1 Practice Test Questions
9. Balancing eBooks and Physical Books Chem 110 Chapter 1 Practice Test Questions
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem 110 Chapter 1 Practice Test Questions
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem 110 Chapter 1 Practice Test Questions
- Setting Reading Goals Chem 110 Chapter 1 Practice Test Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem 110 Chapter 1 Practice Test Questions

- Fact-Checking eBook Content of Chem 110 Chapter 1 Practice Test Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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