

Name Key

1. Hydrogen,  $H_2$ , is used as rocket fuel. The hydrogen is burned in oxygen,  $O_2$ , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen,  $H_2$  is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

$$99.2 \text{ mol } H_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol } H_2} = -24,006 \text{ kJ}$$

-24,006 kJ  
~~24,006 kJ~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is  $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$ . How much heat must be added to 5.2 g of iron at  $25^\circ\text{C}$  to raise its temperature to  $225^\circ\text{C}$ ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine,  $N_2H_4$ , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



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**BM King**



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### **Table of Contents Chemistry Practice Problems With Answers Activity Series**

1. Understanding the eBook Chemistry Practice Problems With Answers Activity Series
  - The Rise of Digital Reading Chemistry Practice Problems With Answers Activity Series
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Practice Problems With Answers Activity Series
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Platform
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Practice Problems With Answers Activity Series
  - Personalized Recommendations
  - Chemistry Practice Problems With Answers Activity Series User Reviews and Ratings
  - Chemistry Practice Problems With Answers Activity Series and Bestseller Lists
5. Accessing Chemistry Practice Problems With Answers Activity Series Free and Paid eBooks
  - Chemistry Practice Problems With Answers Activity Series Public Domain eBooks
  - Chemistry Practice Problems With Answers Activity Series eBook Subscription Services

- Chemistry Practice Problems With Answers Activity Series Budget-Friendly Options
6. Navigating Chemistry Practice Problems With Answers Activity Series eBook Formats
    - ePub, PDF, MOBI, and More
    - Chemistry Practice Problems With Answers Activity Series Compatibility with Devices
    - Chemistry Practice Problems With Answers Activity Series Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Practice Problems With Answers Activity Series
    - Highlighting and Note-Taking Chemistry Practice Problems With Answers Activity Series
    - Interactive Elements Chemistry Practice Problems With Answers Activity Series
  8. Staying Engaged with Chemistry Practice Problems With Answers Activity Series
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Practice Problems With Answers Activity Series
  9. Balancing eBooks and Physical Books Chemistry Practice Problems With Answers Activity Series
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Practice Problems With Answers Activity Series
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Practice Problems With Answers Activity Series
    - Setting Reading Goals Chemistry Practice Problems With Answers Activity Series
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Practice Problems With Answers Activity Series
    - Fact-Checking eBook Content of Chemistry Practice Problems With Answers Activity Series
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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