

$$K_c = \frac{0.6 + x}{(0.2 - x)(0.3 - x)} = 30$$

$$0.6 + x = 30(0.06 - 0.2x - 0.3x + x^2)$$

$$0.6 + x = 1.8 - 15x + 30x^2$$

$$30x^2 - 16x + 1.2 = 0$$

$$x_{1,2} = \frac{16 \pm \sqrt{256 - 144}}{60} = \frac{16 \pm 10.6}{60}$$

$$x_1 = 0.44 \quad x_2 = 0.09$$

$$[SO_2] = 0.2 - 0.09 = 0.011 \text{ mol/l}$$

$$[Cl_2] = 0.3 - 0.09 = 0.021 \text{ mol/l}$$

$$[SO_2 Cl_2] = 0.6 - 0.09 = 0.69 \text{ mol/l}$$

3) Per il seguente equilibrio in fase omogenea la K_c vale 5.06:



Sapendo che ad un certo stadio del processo le concentrazioni istantanee di CO, H₂O, CO₂, e H₂ sono rispettivamente 0.1 M, 0.2 M, 1 M e 0.5M, stabilire se la reazione è all'equilibrio e, in caso contrario, dire in che senso avverrà e calcolare la composizione all'equilibrio (si consideri il volume di 1 litro).

| | | | | | | | |
|--------|-----|---|------------------|---|-----------------|---|----------------|
| | CO | + | H ₂ O | ⇌ | CO ₂ | + | H ₂ |
| inizio | 0.1 | | 0.2 | | 1 | | 0.5 |

$$Q = \frac{[CO_2][H_2]}{[CO][H_2O]} = \frac{1 \cdot 0.5}{0.1 \cdot 0.2} = 25 \quad Q > K_c \quad \text{quindi la reazione è spostata verso sinistra}$$



| | | | | | | | |
|------------|---------|---|------------------|---|-----------------|---|----------------|
| | CO | + | H ₂ O | ⇌ | CO ₂ | + | H ₂ |
| inizio | 0.1 | | 0.2 | | 1 | | 0.5 |
| equilibrio | (0.1+x) | | (0.2+x) | | (1-x) | | (0.5-x) |

$$K_c = \frac{[CO_2][H_2]}{[CO][H_2O]}$$

$$5.06 = \frac{(1-x)(0.5-x)}{(0.1+x)(0.2+x)} \quad 5.06 = \frac{0.5-x-0.5x+x^2}{0.02+0.1x+0.2x+x^2} \quad 5.06 = \frac{0.5-1.5x+x^2}{0.02+0.3x+x^2}$$

Equilibrio Chimico Esercizi Svolti Chemicamo Org

JG Myers



Equilibrio Chimico Esercizi Svolti Chemicamo Org:

Esercizi svolti sull'equilibrio chimico Alessandro Ceriotti, Francesca Porta, 2007 **Chimica generale: esercizi svolti** Stefania Lamponi, **Esercizi svolti. Chimica generale. Principi ed applicazioni moderne**, 2015 **Esercizi di chimica** Silvia Ronchetti, Barbara Onida, Alessandro Delmastro, 2019-11-01 Questo nuovo libro di esercizi di Chimica vuole essere un utile strumento didattico indirizzato in modo particolare alla preparazione dell'esame scritto dei corsi di Chimica del primo anno di Ingegneria Il testo organizzato in sette capitoli che riportano esercizi di stechiometria risolti attraverso procedimenti commentati e accompagnati da note All'inizio di ciascuno lo studente troverà alcuni brevi richiami teorici il cui taglio si riferisce espressamente ai corsi di Chimica per Ingegneria Il percorso didattico comincia affrontando sinteticamente il bilanciamento delle reazioni chimiche per proseguire con i rapporti ponderali tra i reagenti e i prodotti Seguono esercizi sullo stato gassoso leggi dei gas e miscele reattive e sulle soluzioni concentrazione e proprietà colligative Il quinto capitolo dedicato alla termochimica con particolare attenzione al calcolo del calore associato alla trasformazione chimica e il sesto all'equilibrio chimico omogeneo ed eterogeneo in fase gas e in soluzione acquosa L'ultimo capitolo di elettrochimica riporta esercizi su sistemi voltaici ed elettrolitici Lo studente potrà infine verificare le conoscenze acquisite attraverso gli esercizi di autovalutazione proposti al termine di ogni capitolo Tutti gli esercizi si basano sui concetti e sulle leggi principali della Chimica Generale e Inorganica Nella prospettiva in cui gli aspetti quantitativi affrontati attraverso il calcolo siano imprescindibili per la conoscenza della Chimica di base il testo si propone come parte integrante degli strumenti didattici forniti durante il corso **Chimica esercizi svolti: moli, composti, formule** Alessio Mangoni, 2017 *Esercizi di chimica* Cosimo F. Nobile, Piero Mastrorilli, 1999 *Esercizi svolti di chimica fisica* Gianfrancesco Berchiesi, Carlo Santini, 1998 *Chimica generale* Jerome L. Rosenberg, 1977 Come risolvere i problemi di chimica. 400 esercizi svolti e 150 esercizi da svolgere Franco Ugozzoli, 2013 Spontaneamente. Esercizi risolti di chimica fisica Pierandrea Lo Nostro, Niccolò Peruzzi, 2013 **Esercizi svolti di chimica** Donato Vanni, Sandro Vanni, 1987 **Chimica. Esercizi** Valeria Balboni, 2010 **Chimica** Fulvio Cacace, Mario Schiavello, 1973 **Formule, moli e reazioni** Alessio Mangoni, 2025-05-27 Questo testo didattico completo e rigoroso dedicato allo studio e all'applicazione dei fondamenti della chimica quantitativa Attraverso sessanta esercizi svolti con chiarezza e precisione affronta temi essenziali quali massa molare e peso molecolare concetto di mole numero di Avogadro calcolo di atomi e molecole in quantità note composizione percentuale dei composti determinazione delle formule minima e molecolare bilanciamento di equazioni chimiche Si propone come uno strumento efficace per studenti di chimica e per chiunque desideri approfondire le basi della disciplina attraverso un approccio pratico e metodico fornendo solide competenze indispensabili per il proseguimento degli studi scientifici

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