

Chemistry - Chapter 9 Assessment Answers (Pages 281-282)

42. a) 2- b) 1+ c) 1- d) 3+

43. a) 2+ b) 2+ c) 3+ d) 1+

44. a) barium ion b) iodide ion c) silver ion d) mercury(II) ion

45. Cyanide, CN^- and hydroxide, OH^-

46. a) hydroxide ion b) lead(IV) ion c) sulfate ion d) oxide ion

47. zero; A compound is electrically neutral.

48. The symbols for the cation and anion are written and the charges are balanced with subscripts. The name of the cation is followed by the name of the anion.

49. Determine the charge of the anion then work the formula backwards to find the charge of the transition metal cation needed to give a net charge of zero for the formula unit.

50. The symbols for the cation and anion are written and the charges are balanced with subscripts. Parentheses are used around the polyatomic ion if a subscript is needed. The name of the cation is followed by the name of the anion.

51. a and b

52. When more than a single polyatomic ion is needed to balance the formula.

53. NH_4NO_3 , ammonium nitrate; $(\text{NH}_4)_2\text{CO}_3$, ammonium carbonate; NH_4CN , ammonium cyanide; $(\text{NH}_4)_3\text{PO}_4$, ammonium phosphate; $\text{Sn}(\text{NO}_3)_4$, tin(IV) nitrate; $\text{Sn}(\text{CO}_3)_2$, tin(IV) carbonate; $\text{Sn}(\text{CN})_4$, tin(IV) cyanide; $\text{Sn}_2(\text{PO}_4)_4$, tin(IV) phosphate; $\text{Fe}(\text{NO}_3)_3$, iron(III) nitrate; $\text{Fe}_2(\text{CO}_3)_3$, iron(III) carbonate; $\text{Fe}(\text{CN})_3$, iron(III) cyanide; FePO_4 , iron(III) phosphate; $\text{Mg}(\text{NO}_3)_2$, magnesium nitrate; MgCO_3 , magnesium carbonate; $\text{Mg}(\text{CN})_2$, magnesium cyanide; $\text{Mg}_3(\text{PO}_4)_2$, magnesium phosphate

54. two nonmetals

55. a) tri- b) mono- c) di- d) hexa- e) penta- f) tetra-

56. For formulas, write the correct symbols for each element with a subscript corresponding to the prefix before each element in the name. For naming, name each element in the order given. Use the subscript to determine the prefixes before each element in the name. The name ends in *-ide*.

57. a) BCl_3 b) dinitrogen pentoxide c) N_2H_4 d) carbon tetrachloride

58. a) hydrochloric acid b) H_2SO_4 c) nitric acid d) $\text{HC}_2\text{H}_3\text{O}_2$

Mastering Chemistry Chapter 9 Answers

DJ Losen



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Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science **Mastering Chemistry** Jason Chin, Vault Reports (Firm), Job Vault, 1998 The periodic table never looked so good From the properties of matter and mole calculations to the kinetics of reactions and enthalpy change this guide demystifies the complexities of chemistry through easy to follow charts graphs and sample problems **EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS** CHANG, 2013-01-07 EBOOK GENERAL CHEMISTRY THE ESSENTIAL CONCEPTS *Mastering Chemistry* Gerald W. Gibson, 1975 *General Chemistry for Engineers* Jeffrey Gaffney, Nancy Marley, 2017-11-13 General Chemistry for Engineers explores the key areas of chemistry needed for engineers This book develops material from the basics to more advanced areas in a systematic fashion As the material is presented case studies relevant to engineering are included that demonstrate the strong link between chemistry and the various areas of engineering Serves as a unique chemistry reference source for professional engineers Provides the chemistry principles required by various engineering disciplines Begins with an atoms first approach building from the simple to the more complex chemical concepts Includes engineering case studies connecting chemical principles to solving actual engineering problems Links chemistry to contemporary issues related to the interface between chemistry and engineering practices **Chemistry** John S. Phillips, Cheryl Wistrom, 2000 **Stewart's Textbook of Acid-Base** John A Kellum, Paul WG Elbers, 2009 Rev ed of How to understand acid base c1981 **A Problem-Solving Approach to Aquatic Chemistry** James N. Jensen, 2022-12-20 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter s main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases

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Modern University Chemistry, Custom Pub Norbert T. Porile, 1993

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