

# Chapter - 1

## Chemical Reactions and Equations

Chemical Reaction: • Chemical reaction is a process in which new substances with different properties are formed.

- Atoms of one element don't change into another element.
- Only rearrangement of atoms takes place in a chemical reaction.

Examples 1. Hydrogen + Oxygen  $\rightarrow$  water

2. Iron + Oxygen  $\rightarrow$  rust

3. Potassium and chlorine gas  $\rightarrow$  chloride

Physical Change: • No new substance is formed.

- It is a temporary change.
- No change in the composition of the substance takes place.
- It is easily reversible.
- Heat or light may or may not be given out or absorbed.

Chemical Change: • A new substance is formed.

- It is a permanent change.
- The composition of the new substance change greatly.
- The change may be reversible or irreversible.
- Heat or light or both are generally given out or absorbed.

## Law of Conservation of mass

Matter can neither be created nor destroyed in a chemical reaction. This means that total mass of all reactants must

# Chemistry Chapter 1 Notes

**J Ma**



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## **Chemistry Chapter 1 Notes** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Chemistry Chapter 1 Notes**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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