

## AP Chemistry: Chapter 14 Practice Test

10. Given the following acids and  $K_a$  values:

$\text{HClO}_4$	$\text{HOAc}$	$\text{HCN}$	$\text{HF}$
-----------------	---------------	--------------	-------------

$1 \times 10^7$	$1.76 \times 10^{-5}$	$4.93 \times 10^{-10}$	$3.53 \times 10^{-4}$
-----------------	-----------------------	------------------------	-----------------------

which shows the conjugate bases listed by increasing strength?

- |  |  |
|--|--|
| [A] $\text{ClO}_4^-$ , $\text{F}^-$ , $\text{OAc}^-$ , $\text{CN}^-$ | [B] $\text{CN}^-$ , $\text{F}^-$ , $\text{OAc}^-$ , $\text{ClO}_4^-$ |
| [C] $\text{CN}^-$ , $\text{OAc}^-$ , $\text{F}^-$ , $\text{ClO}_4^-$ | [D] $\text{CN}^-$ , $\text{ClO}_4^-$ , $\text{F}^-$ , $\text{OAc}^-$ |
| [E] $\text{ClO}_4^-$ , $\text{OAc}^-$ , $\text{CN}^-$ , $\text{F}^-$ |  |
- 
15. Given that the  $K_a$  for  $\text{HOCl}$  is  $3.5 \times 10^{-8}$ , calculate the  $K$  value for the reaction of  $\text{HOCl}$  with  $\text{OH}^-$ .
- |                           |                       |                          |
|---------------------------|-----------------------|--------------------------|
| [A] $3.5 \times 10^{-22}$ | [B] $3.5 \times 10^6$ | [C] $2.9 \times 10^{-7}$ |
| [D] $3.5 \times 10^{-8}$  | [E] none of these     |                          |
- 
32. Acetic acid ( $\text{HC}_2\text{H}_3\text{O}_2$ ) is a weak acid ( $K_a = 1.8 \times 10^{-5}$ ). Calculate the pH of a 17.6 M  $\text{HC}_2\text{H}_3\text{O}_2$  solution.
- |         |         |         |          |         |
|---------|---------|---------|----------|---------|
| [A] 4.3 | [B] 6.4 | [C] 1.7 | [D] 0.97 | [E] 7.4 |
|---------|---------|---------|----------|---------|
- 
52. What concentration of acetic acid ( $K_a = 1.80 \times 10^{-5}$ ) has the same pH as that of  $5.00 \times 10^{-3}$  M  $\text{HCl}$ ?
- |                             |                             |            |
|-----------------------------|-----------------------------|------------|
| [A] $5.00 \times 10^{-3}$ M | [B] $3.60 \times 10^{-3}$ M | [C] 1.39 M |
| [D] 5.00 M                  | [E] none of these           |            |

Consider the reaction  $\text{HOCl} + \text{F}^- \rightleftharpoons \text{HF} + \text{OCl}^-$

# Chemistry Chapter 14 Test

**K Morrison**



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