

Project 1: Topic and system (or product) concept statement

Overview

In this project assignment you get started by establishing the target product or system that you will work on for a client and by writing a system (or product) concept statement.

What To Do

1. Establish a client for your project (virtual or fictitious is ok).
2. Establish a product or application system for which you will develop the user interaction design.
3. Get your client and target product or application system approved by the instructor before the due date for this project part. Earlier approval is better, in case you have to make some adjustments or find a different client.
4. Write and refine (several times) a system concept statement for your target system.

The Importance of Making These Choices

We urge you to choose carefully, since you will be living with your decision for the whole semester. Your choices for a product or system can range from an information system to a Web site to a commercial product such as a new kind of mobile or personal device. You can even design a new system with gestural interaction on a wall-size display. We are requiring you to pick a domain such as those discussed in class on the first day of lectures and listed below. These are not requirements but

- Car + Phone for music experience, or navigation (GPS), or shopping with location based reminders, etc.
- Large public displays in common areas coupled with your phones and proximity
- Bike + exercise coupled with your iDevice to manage your health, fitness, competitive spirit
- Public transportation + iDevice for daily travel or vacations, save energy, bike, car, share-ride, etc.

Please consider a choice that will involve lots of creative design, as that is one of the important emphases in this course. We also urge you to consider a product

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Robert R. Redfield



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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$
 $8.3 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N? what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO}_2 + 3\text{H}_2\text{O}$ if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g } \text{O}_2/\text{mole}} = 0.666 \text{ moles } \text{O}_2$
 $\frac{21.3 \text{ g } \text{C}_3\text{H}_6}{42 \text{ g } \text{C}_3\text{H}_6/\text{mole}} = 0.507 \text{ moles } \text{C}_3\text{H}_6$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
b determine the theoretical yield of Al_2O_3
 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3/\text{mol} = 302.3 \text{ g } \text{Al}_2\text{O}_3$
c determine the percent yield
 $\frac{260.0 \text{ g } \text{Al}_2\text{O}_3}{302.3 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 86.01\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H_2 are left over 13

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